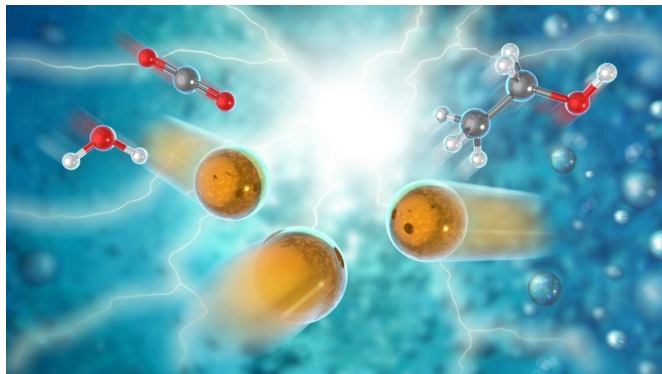


Researchers discover new electrocatalyst for turning carbon dioxide into liquid fuel

5 August 2020, by Joseph E. Harmon



Artistic rendering of electrocatalytic process for conversion of carbon dioxide and water into ethanol. Credit: Argonne National Laboratory

Catalysts speed up chemical reactions and form the backbone of many industrial processes. For example, they are essential in transforming heavy oil into gasoline or jet fuel. Today, catalysts are involved in over 80 percent of all manufactured products.

A research team, led by the U.S. Department of Energy's (DOE) Argonne National Laboratory in collaboration with Northern Illinois University, has discovered a new electrocatalyst that converts carbon dioxide (CO₂) and water into ethanol with very high energy efficiency, high selectivity for the desired final product and low cost. Ethanol is a particularly desirable commodity because it is an ingredient in nearly all U.S. gasoline and is widely used as an intermediate product in the chemical, pharmaceutical and cosmetics industries.

"The process resulting from our [catalyst](#) would contribute to the circular carbon economy, which entails the reuse of carbon dioxide," said Di-Jia Liu, senior chemist in Argonne's Chemical Sciences and Engineering division and a UChicago CASE scientist in the Pritzker School of Molecular

Engineering, University of Chicago. This process would do so by electrochemically converting the CO₂ emitted from [industrial processes](#), such as fossil fuel power plants or alcohol fermentation plants, into valuable commodities at reasonable cost.

The team's catalyst consists of atomically dispersed copper on a carbon-powder support. By an electrochemical reaction, this catalyst breaks down CO₂ and [water molecules](#) and selectively reassembles the broken molecules into ethanol under an external electric field. The electrocatalytic selectivity, or "Faradaic efficiency," of the process is over 90 percent, much higher than any other reported process. What is more, the catalyst operates stably over extended operation at low voltage.

"With this research, we've discovered a new catalytic mechanism for converting carbon dioxide and water into ethanol," said Tao Xu, a professor in physical chemistry and nanotechnology from Northern Illinois University. "The mechanism should also provide a foundation for development of highly efficient electrocatalysts for [carbon dioxide](#) conversion to a vast array of value-added chemicals."

Because CO₂ is a stable molecule, transforming it into a different molecule is normally energy intensive and costly. However, according to Liu, "We could couple the electrochemical process of CO₂-to-ethanol conversion using our catalyst to the electric grid and take advantage of the low-cost electricity available from renewable sources like solar and wind during off-peak hours." Because the process runs at low temperature and pressure, it can start and stop rapidly in response to the intermittent supply of the renewable electricity.

The team's research benefited from two DOE Office of Science User Facilities at Argonne—the Advanced Photon Source (APS) and Center for

Nanoscale Materials (CNM)—as well as Argonne's Laboratory Computing Resource Center (LCRC). "Thanks to the high photon flux of the X-ray beams at the APS, we have captured the structural changes of the catalyst during the [electrochemical reaction](#)," said Tao Li, an assistant professor in the Department of Chemistry and Biochemistry at Northern Illinois University and an assistant scientist in Argonne's X-ray Science division. These data along with high-resolution electron microscopy at CNM and computational modeling using the LCRC revealed a reversible transformation from atomically dispersed copper to clusters of three copper atoms each on application of a low voltage. The CO₂-to-ethanol catalysis occurs on these tiny copper clusters. This finding is shedding light on ways to further improve the catalyst through rational design.

"We have prepared several new catalysts using this approach and found that they are all highly efficient in converting CO₂ to other hydrocarbons," said Liu. "We plan to continue this research in collaboration with industry to advance this promising technology."

More information: Haiping Xu et al, Highly selective electrocatalytic CO₂ reduction to ethanol by metallic clusters dynamically formed from atomically dispersed copper, *Nature Energy* (2020). [DOI: 10.1038/s41560-020-0666-x](https://doi.org/10.1038/s41560-020-0666-x)

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