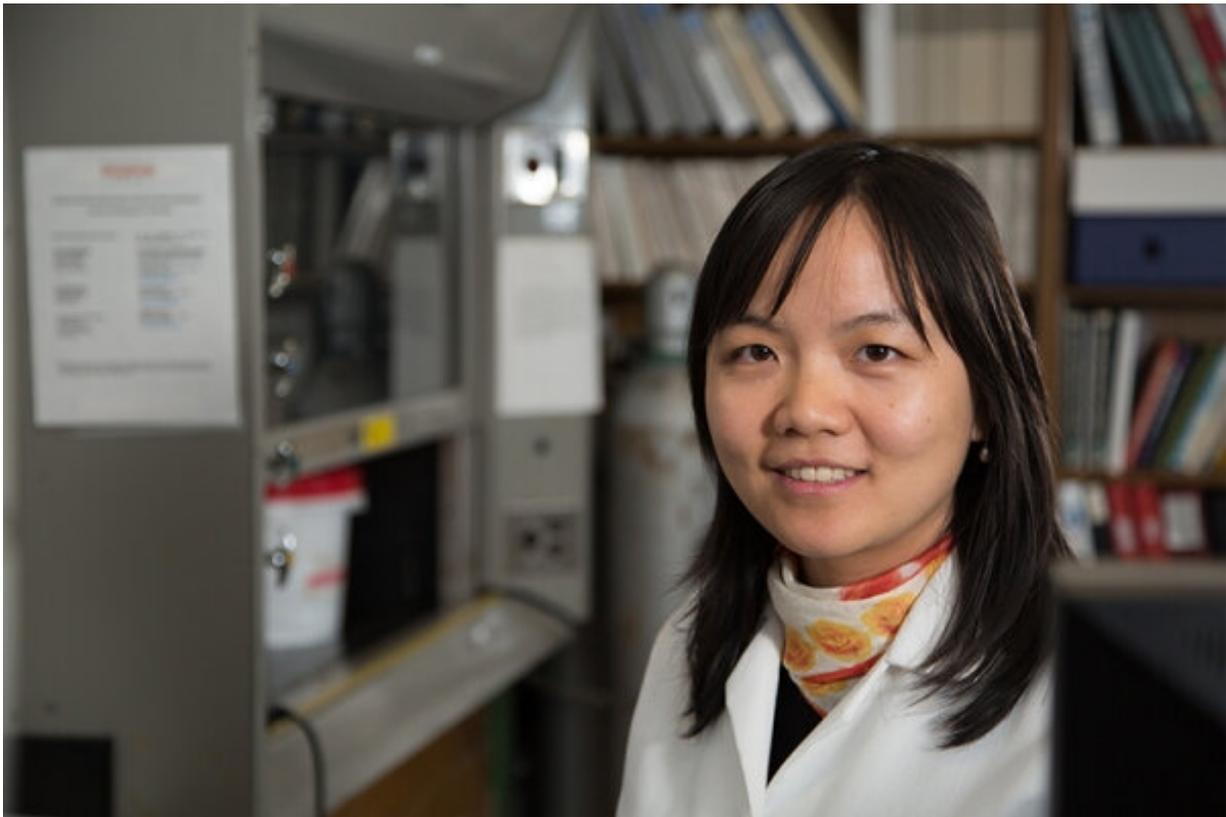


# Solving a scientific mystery and finding a solution for industry

May 21 2019

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Yandi Hu, assistant professor of civil and environmental engineering at the University of Houston, led a team of researchers in developing a better understanding of the presence of strontium-rich barite in seawater. Credit: University of Houston

In solving a scientific mystery, researchers from the University of

Houston and the nation's national laboratories also discovered a new avenue for clearing toxins from water, including wastewater produced by hydraulic fracturing.

The findings, described in a paper published this week in the *Proceedings of the National Academy of Sciences*, address the paradox of why the mineral barite—often laced with the chemical element [strontium](#)—is widely present in seawater.

Seawater is undersaturated with respect to barite, that is, the concentrations of barium and sulfate ions are too low to form barite precipitates, said Yandi Hu, assistant professor of civil and [environmental engineering](#) at the University of Houston and corresponding author on the paper. Strontium sulfate is much more water-soluble than barium sulfate, so thermodynamic levels of strontium in barium sulfate precipitates should be low, as well, she said.

But barite is often found in seawater, as is strontium-laced barite. Based on field-site findings that the presence of barite and high strontium levels in barite are associated with the presence of organic matter, she and colleagues began their study.

In addition to Hu, co-authors for the paper include Ning Deng and Bo Cao of UH; Andrew Stack and Julian Weber of the Oak Ridge National Laboratory; and James De Yoreo of the Pacific Northwest National Laboratory.

"An understanding of barite precipitation in the ocean, which is globally undersaturated with respect to barite, is missing," they wrote.

"Moreover, the reason for the occurrence of higher [strontium] content in marine barites than expected... remains unknown."

Field data suggested barite precipitates were more likely to be found in

seawater if [organic material](#) was nearby, although the researchers said the specific mechanism for mineral-organic interactions are not clear.

To better understand why, they used organic films in both undersaturated and supersaturated solutions and recorded barite nucleation. Tests found barite grew quickly in the solution, with continuous nucleation of barite at the interface between the solution and the organic film.

That interface creates conditions—a microenvironment—that is different from that of the main body of the solution, Hu said.

"The organics can enrich ions from the bulk solution, making the local solution at the interface supersaturated with respect to barite, even when the bulk solution was undersaturated," Hu said.

The supersaturation along the organic film promoted continuous nucleation of barite there, while the fact that the strontium sulfate exhibited a lower nucleation energy prompted the incorporation of strontium into the precipitates, she said. "That explains the mysteries."

In addition to offering an explanation for strontium-rich barite formation in marine environments, the researchers said the work offers new insights for understanding and manipulating solid solution nucleation and growth, which could lead to new methods for removing toxins from water and other liquids. That includes the possibility of a more efficient way to remove strontium from the produced water resulting from hydraulic fracturing, Hu said.

She said both barite and strontium are often found in the produced water; while it is relatively easy to remove barite with the addition of a [sulfate](#), removing strontium is more complicated. "We propose that with an organic presence, you can remove strontium more effectively,"

**More information:** Ning Deng et al., "Organic–mineral interfacial chemistry drives heterogeneous nucleation of Sr-rich  $(\text{Ba}_x, \text{Sr}_{1-x})\text{SO}_4$  from undersaturated solution," *PNAS* (2019).

[www.pnas.org/cgi/doi/10.1073/pnas.1821065116](http://www.pnas.org/cgi/doi/10.1073/pnas.1821065116)

Provided by University of Houston

Citation: Solving a scientific mystery and finding a solution for industry (2019, May 21)  
retrieved 16 April 2024 from

<https://phys.org/news/2019-05-scientific-mystery-solution-industry.html>

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