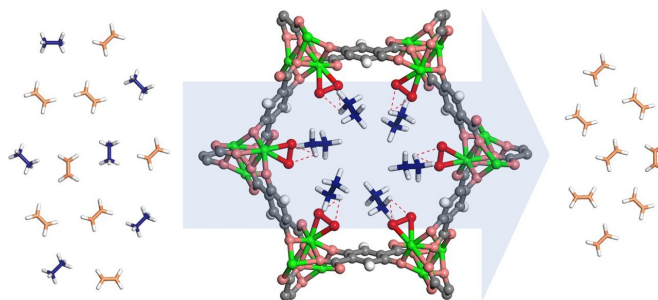


Novel material could make plastic manufacturing more energy-efficient

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This iron-based metal-organic framework decorated with peroxy groups (red and green) can capture ethane (blue molecules) while allowing ethylene (peach molecules) to pass through, potentially providing a more efficient and cost-effective way to purify ethylene, the most important raw material for plastic production. Credit: Wei Zhou / NIST

An innovative filtering material may soon reduce the environmental cost of manufacturing plastic. Created by a team including scientists at the National Institute of Standards and Technology (NIST), the advance can extract the key ingredient in the most common form of plastic from a mixture of other chemicals—while consuming far less energy than usual.

The material is a metal-organic framework (MOF), a class of substances that have repeatedly demonstrated a talent for separating individual hydrocarbons from the soup of organic molecules produced by [oil refining](#) processes. MOFs hold immense value for the plastic and petroleum industries because of this capability, which could allow manufacturers to perform these separations far more cheaply than standard oil-refinement techniques.

This promise has made MOFs the subject of intense study at NIST and elsewhere, leading to MOFs that can separate different octanes of

gasoline and speed up complex chemical reactions. One major goal has proved elusive, though: an industrially preferred method for wringing out ethylene—the molecule needed to create polyethylene, the plastic used to make shopping bags and other everyday containers.

However, in today's issue of the journal *Science*, the research team reveals that a modification to a well-studied MOF enables it to separate purified ethylene out of a mixture with ethane. The team's creation—built at The University of Texas at San Antonio (UTSA) and China's Taiyuan University of Technology and studied at the NIST Center for Neutron Research (NCNR)—represents a major step forward for the field.

Making plastic takes lots of energy. Polyethylene, the most common type of plastic, is built from ethylene, one of the many [hydrocarbon](#) molecules found in crude oil refining. The ethylene must be highly purified for the manufacturing process to work, but the current industrial technology for separating ethylene from all the other hydrocarbons is a chilly but high-energy process that cools down the crude to more than 100 degrees below zero Celsius.

Ethylene and ethane constitute the bulk of the hydrocarbons in the mixture, and separating these two is by far the most energy-intensive step. Finding an alternative method of separation would reduce the energy needed to make the 170 million tons of ethylene manufactured worldwide each year.

Scientists have been searching for such an alternative method for years, and MOFs appear promising. On a microscopic level, they look a bit like a half-built skyscraper of girders and no walls. The girders have surfaces that certain [hydrocarbon molecules](#) will stick to firmly, so pouring a mixture of two hydrocarbons through the right MOF can pull one kind of molecule out of the mix, letting the other

hydrocarbon emerge in pure form.

The trick is to create a MOF that allows the ethylene to pass through. For the plastics industry, this has been the sticking point.

"It's very difficult to do," said Wei Zhou, a scientist at the NCNR. "Most MOFs that have been studied grab onto ethylene rather than ethane. A few of them have even demonstrated excellent separation

performance, by selectively adsorbing the ethylene. But from an industrial perspective you would prefer to do the opposite if feasible. You want to adsorb the ethane byproduct and let the ethylene pass through."

The research team spent years trying to crack the problem. In 2012, another research team that worked at the NCNR found that a particular framework called MOF-74 was good for separating a variety of hydrocarbons, including [ethylene](#). It seemed like a good starting point, and the team members scoured the scientific literature for additional inspiration. An idea taken from biochemistry finally sent them in the right direction.

"A huge topic in chemistry is finding ways to break the strong bond that forms between carbon and hydrogen," said UTSA professor Banglin Chen, who led the team. "Doing that allows you to create a lot of valuable new materials. We found previous research that showed that compounds containing iron peroxide could break that bond."

The team reasoned that to break the bond in a hydrocarbon molecule, the compound would have to attract the molecule in the first place. When they modified MOF-74's walls to contain a structure similar to the compound, it turned out the molecule it attracted from their mixture was ethane.

The team brought the MOF to the NCNR to explore its atomic structure. Using a technique called neutron diffraction, they determined what part of the MOF's surface attracts ethane—a key piece of information for explaining why their innovation succeeded where other efforts have fallen short.

"Without the fundamental understanding of the

mechanism, no one would believe our results," Chen said. "We also think that we can try to add other small groups to the surface, maybe do other things. It's a whole new research direction and we're very excited."

While Zhou said the team's modified MOF does work efficiently, it may require some additional development to see action at a refinery.

"We proved this route is promising," Zhou said, "but we're not claiming our materials perform so well they can't be improved. Our future goal is to dramatically increase their selectivity. It's worth pursuing further."

More information: L. Li et al., "Ethane/ethylene separation in a metal-organic framework with iron-peroxo sites," *Science* (2018). [science.sciencemag.org/cgi/doi ... 1126/science.aat0586](https://science.sciencemag.org/cgi/doi/10.1126/science.aat0586)

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