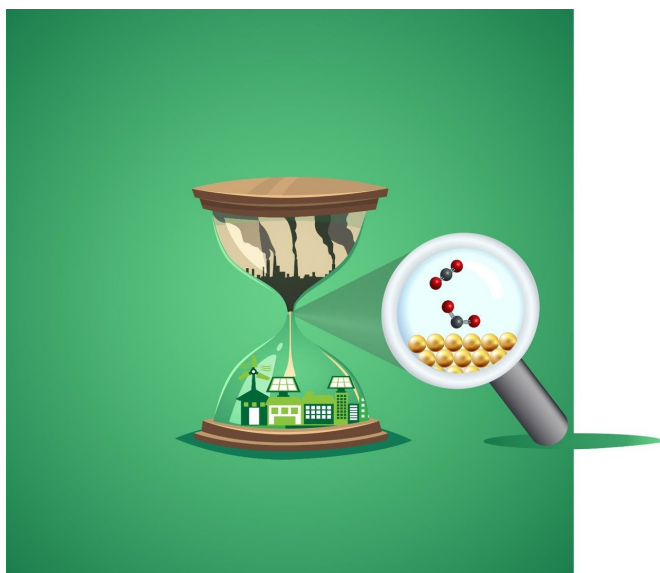


# After 150 years, a breakthrough in understanding the conversion of CO<sub>2</sub> to electrofuels

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Visualization of a shift from air-polluting economy based on fossil fuels to clean economy based on renewable energy, which is facilitated by electrocatalytic conversion of abundant CO<sub>2</sub> to fuels and other useful chemicals. The bottleneck of this reaction is activation of a linear CO<sub>2</sub> molecule to adsorbed carboxylate CO<sub>2</sub><sup>-</sup> with a chair-like geometry. Credit: Irina Chernyshova & Sathish Ponnurangam/Columbia Engineering

Using surface-enhanced Raman spectroscopy, Columbia Engineers are first to observe how CO<sub>2</sub> is activated at the electrode-electrolyte interface; their finding shifts the catalyst design from trial-and-error paradigm to a rational approach and could lead to alternative, cheaper, and safer renewable energy storage

New York, NY—September 17, 2018—Scientists have long sought ways to convert abundant CO<sub>2</sub> to useful products such as chemicals and fuel. As early as in 1869, they were able to electrocatalytically convert CO

to formic acid. Over the past two decades, the rise of CO<sub>2</sub> in the Earth's atmosphere has significantly accelerated research in CO<sub>2</sub> conversion using [renewable energy resources](#), including solar, wind, and tidal. Because these resources are intermittent—the sun doesn't shine every day, nor does the wind constantly blow—how to store renewable [energy](#) safely and cost-effectively is a major challenge.

Recent research in electrocatalytic CO<sub>2</sub> conversion points the way to using CO<sub>2</sub> as a feedstock and renewable electricity as an energy supply for the synthesis of different types of fuel and value-added chemicals such as ethylene, ethanol, and propane. But scientists still do not understand even the first step of these reactions—CO<sub>2</sub> activation, or the transformation of the linear CO<sub>2</sub> molecule at the catalyst surface upon accepting the first electron. Knowing the exact structure of the activated CO<sub>2</sub> is essential because its structure dictates both the end product of the reaction and its energy cost. This reaction can start from many initial steps and go through many pathways, giving typically a mixture of products. If scientists figure out how the process works, they will be better able to selectively promote or inhibit certain pathways, which will lead to the development of a commercially viable catalyst for this technology.

Columbia Engineering researchers announced today that they solved the first piece of the puzzle—they have proved that CO<sub>2</sub> electroreduction begins with one common intermediate, not two as was commonly thought. They applied a comprehensive suite of experimental and theoretical methods to identify the structure of the first intermediate of CO<sub>2</sub> electroreduction: carboxylate CO<sub>2</sub><sup>-</sup>—that is attached to the surface with C and O atoms. Their breakthrough, published online today in *PNAS*, came by applying surface enhanced Raman

scattering (SERS) instead of the more frequently used surface enhanced infrared spectroscopy (SEIRAS). The spectroscopic results were corroborated by quantum chemical modeling.

"Our findings about CO<sub>2</sub> activation will open the door to an incredibly broad range of possibilities: if we can fully understand CO<sub>2</sub> electroreduction, we'll be able to reduce our dependence on fossil fuels, contributing to the mitigation of climate change," says the paper's lead author Irina Chernyshova, associate research scientist, department of earth and environmental engineering. "In addition, our insight into CO<sub>2</sub> activation at the solid-water interface will enable researchers to better model the prebiotic scenarios from CO<sub>2</sub> to complex organic molecules that may have led to the origin of life on our planet."

They decided to use SERS rather than SEIRAS for their observations because they found that SERS has several significant advantages that enable more accurate identification of the structure of the reaction intermediate. Most importantly, the researchers were able to measure the vibrational spectra of species formed at the electrode-electrolyte interface along the entire spectral range and on an operating electrode (in operando). Using both quantum chemical simulations and conventional electrochemical methods, the researchers were able to get the first detailed look at how CO<sub>2</sub> is activated at the electrode-electrolyte interface.

Understanding the nature of the first reaction intermediate is a critical step toward commercialization of the electrocatalytic CO<sub>2</sub> conversion to useful chemicals. It creates a solid foundation for moving away from the trial-and-error paradigm to rational catalyst design. "With this knowledge and computational power," says the paper's co-author Sathish Ponnurangam, a former graduate student and postdoc in Somasundaran's lab who is now an assistant professor of chemical and petroleum engineering at the University of Calgary, Canada, "researchers will be able to predict more accurately the reaction on different catalysts and specify the most promising ones, which can further be synthesized and tested."

"The Columbia Engineering experiments provide such detail that we should be able to obtain very definitive validation of the computational models," says William Goddard, Charles and Mary Ferkel Professor of Chemistry, Materials Science, and Applied Physics at CalTech, who was not involved with the study. "I expect that together with our theory, the Columbia Engineering experiments will provide precise mechanisms to be established and that examining how the mechanisms change for different alloys, surface structures, electrolytes, additives, should enable optimization of the electrocatalysts for water splitting (solar fuels), CO<sub>2</sub> reduction to fuels and organic feedstocks, N<sub>2</sub> reduction to NH<sub>3</sub> to obtain much less expensive fertilizers, all the key problems facing society to obtain the energy and food to accommodate our exploding population."

Electrocatalysis and photocatalysis (the so-called artificial photosynthesis) are among the most promising ways to achieve effective storage for renewable energy. CO<sub>2</sub> electroreduction has been capturing the imagination of researchers for more than 150 years because of its similarity to photosynthesis. Just as a plant converts sunlight into chemical energy, a catalyst converts electrons supplied by renewable energy into chemical energy that is stored in reduced products of CO<sub>2</sub>. In addition to its application for [renewable energy](#), electrocatalysis technology may also enable manned Mars missions and colonies by providing fuel for the return trip and carbonaceous chemicals from the CO<sub>2</sub> that makes up 95 percent of that planet's atmosphere.

"We expect our findings and methodology will spur work on how to make go faster and at a lower energy cost not only electrocatalytic but also photocatalytic CO<sub>2</sub> reduction," says Ponisseril Somasundaran, LaVon Duddleson Krumb Professor of Mineral Engineering, Department of Earth and Environmental Engineering. "In the latter case, a catalyst reduces CO<sub>2</sub> using direct sun light. Even though these two approaches are experimentally different they are microscopically similar—both start with activation of CO<sub>2</sub> upon electron transfer from a catalyst surface. At this point, I believe that both these approaches will dominate the future."

The team is now working to uncover the subsequent reaction steps—to see how CO<sub>2</sub> is further transformed—and to develop superior catalysts based on earth-abundant elements such as Cu (copper) and Sn (tin).

The study is titled "On the origin of the elusive first intermediate of CO<sub>2</sub> electroreduction."

**More information:** Irina V. Chernyshova et al., "On the origin of the elusive first intermediate of CO<sub>2</sub> electroreduction," *PNAS* (2018).

[www.pnas.org/cgi/doi/10.1073/pnas.1802256115](http://www.pnas.org/cgi/doi/10.1073/pnas.1802256115)

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