

A clean way to extract rare earth metals

June 20 2016, by Leah Burrows



Rare earth metals are used in everything from wind turbines to cell phones, electric cars and precision weapons. Credit: Wikicommons

Rare earth metals—those 17 chemically similar elements at the bottom of the periodic table—are in almost every piece of technology we use from cell phones to wind turbines to electric cars. Because these elements are so similar to each other, the process of separating them is time consuming, expensive, and dangerous. Processing one ton of rare earths can produce 2,000 tons of toxic waste.

Now, researchers from the Harvard John A. Paulson School of

Engineering and Applied Sciences (SEAS) may have found a clean alternative.

David Clarke, the Extended Tarr Family Professor of Materials at SEAS, and his graduate student William Bonificio, have developed a method to separate [rare earths](#) using bacteria filters and solutions with pH no lower than hydrochloric acid.

The method was recently described in the journal *Environmental Science and Technology Letters*.

Harvard's Office of Technology Development has filed patents and is actively pursuing commercial opportunities.

"Rare earths are very desirable but only if you can separate them," said Clarke. "It's a big 'if' because how can you separate ions that are almost the same size and the same charge?"

Current separation methods involve hundreds of steps and lots of hazardous chemicals.

Clarke and Bonificio found inspiration in the living world of bacteria. Bacteria filters have long been used to bioabsorb toxic elements from wastewater or filter metals from mine drainage systems. Recent research showed that some rare earths can bioabsorb, but Clarke and Bonificio wondered if all rare earths could be filtered through bacteria.

They immobilized a bacteria from marine algae on an assay filter and passed a solution of mixed rare earths (known as lanthanides) through it. The bacteria bioabsorbed all the elements as they passed—plucking them out of the solution and fixing them to their surface.

Then the team pumped solutions of various pH balances through the

filter. With each successive pH wash, different rare earths detached. The researchers found that lighter lanthanides, such as Europium and Praseodymium, desorbed with higher-pH washes while heavier lanthanides, such as Thulium, Lutetium, and Ytterbium, desorbed with lower pH.

The team also found that if they wanted to separate only the heaviest metals, such as Thulium, which is commonly used in lasers and portable X-rays, they could block the bacteria's receptors that absorb the lighter rare earths and only use a low-pH solution.

"We found that it is possible to concentrate a solution of equal concentrations of each lanthanide to nearly 50 percent of the three heaviest lanthanides in just two passes," said Bonificio. "This surpasses existing industrial practice."

"This is a radically different way of doing separation," said Clarke. "We have an opportunity to harness the diversity of bacterial surface chemistry to separate and recover these valuable metals in a way that is environmentally benign."

More information: William D. Bonificio et al. Rare-Earth Separation Using Bacteria, *Environmental Science & Technology Letters* (2016).

[DOI: 10.1021/acs.estlett.6b00064](https://doi.org/10.1021/acs.estlett.6b00064)

Provided by Harvard University

Citation: A clean way to extract rare earth metals (2016, June 20) retrieved 19 May 2024 from <https://phys.org/news/2016-06-rare-earth-metals.html>

This document is subject to copyright. Apart from any fair dealing for the purpose of private

study or research, no part may be reproduced without the written permission. The content is provided for information purposes only.