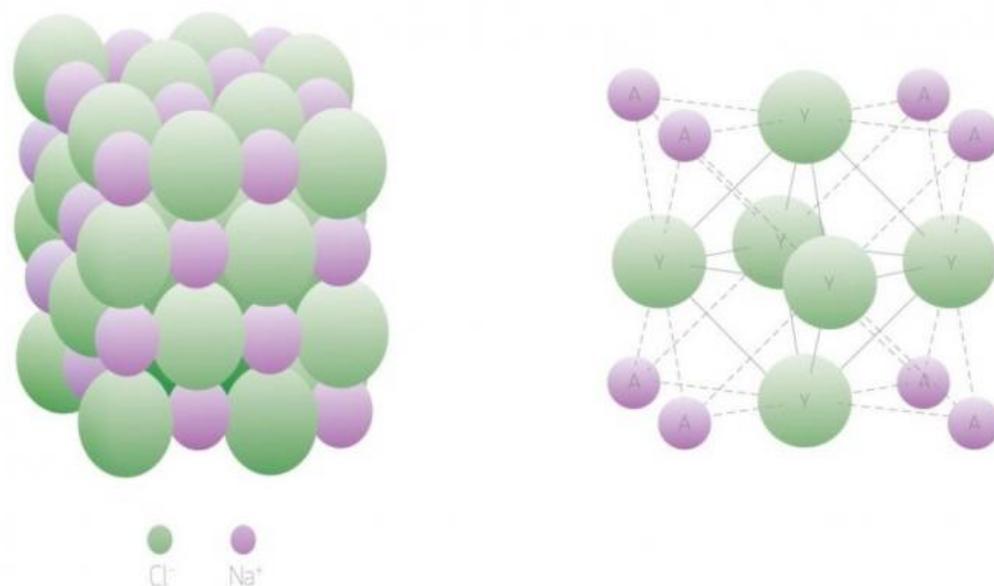


Scientists gain insights into 'forbidden' chemistry

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The structures of 'textbook' NaCl (left) and 'forbidden' A₃Y (A=Li, Na, K; Y=Cl, Br) (right) are shown. Credit: MIPT press office

Gabriele Saleh, a research fellow at MIPT, and Prof. Artem Oganov, a Laboratory Supervisor at MIPT and Professor at the Skolkovo Institute of Science and Technology (Skoltech), have discovered what causes the stability of various compounds that are not commonly found in

'textbook' chemistry.

The reorganisation of the chemical [interactions](#) results in the [stability](#) of the 'new' structure of the compounds. The results of the study have been published in the journal *Physical Chemistry & Chemical Physics*.

Oganov and Saleh derived a simple model and formulated the basic principles of stability of 'forbidden' substances. In addition, the scientists updated the phase diagram of compounds formed by Na and Cl: They added one new compound, Na_4Cl_3 , and two new phases of Na_3Cl . The research was conducted using the USPEX algorithm combined with quantum mechanical calculations.

"We showed how the insights gained in the present study can be used to rationalize the stability of recently discovered high-pressure compounds," say the authors in their paper.

In a previous experiment, Oganov and his colleagues discovered several 'forbidden' compounds- Na_3Cl , NaCl_3 , NaCl_7 , and even Na_3Cl_2 . These compounds are only stable under extreme pressure (approximately 200,000 atmospheres) and they decompose under normal conditions on Earth. However, understanding how new compounds become stable under high pressure is of utmost importance for planetary science.

The principle that explains the unusual ratio of Na and Cl atoms in 'forbidden' compounds is that the number of interactions between Na and Cl atoms increases while interactions between sodium atoms break down.

The interactions between neighbouring atoms in a crystal are responsible for the structure and properties of the crystal (remember carbon and graphite).

Upon formation of these 'forbidden compounds', new Na-Cl interactions are formed at the expenses of Na-Na metallic bonds. The competition between these two bonding types, influenced by pressure, determines the peculiar structures of the newly discovered compounds.

In addition to explaining the stability of sodium subchlorides, Saleh and Oganov suggested that Na could be replaced by any alkali metal, and Cl could be replaced by any heavier halogen. As a result, the model can be used to study the properties of alkali subhalides and predict their stability.

The formation of alkali subhalides, and the mechanism responsible for it, were investigated by considering the reaction: $2\text{Na} + \text{NaCl} \rightarrow \text{Na}_3\text{Cl}$. Or more generally: $2\text{A} + \text{AY} \rightarrow \text{A}_3\text{Y}$ (A=Li, Na, K; Y= F, Cl, Br). Along this reaction, additional bonds and interactions are formed and the coordination number (that is the number of interactions a given atom forms) of halogens increases.

The calculations made by Oganov and Saleh have demonstrated that in addition to the compounds discovered in 2013, Na_4Cl_3 is also stable, and Na_3Cl has two new structures. The investigation of the stability of substances formed along the reaction $2\text{A} + \text{AY} \rightarrow \text{A}_3\text{Y}$ (A=Li, Na, K; Y= F, Cl, Br) led the authors to predict that Li_3Cl , Li_3Br , and Na_3Br are stable under pressure. All three of these subhalides have a structure similar to the structure of Na_3Cl , which was discovered previously.

Each rule must have its exceptions— K^3Br and K_3Cl , for example, display completely different structures. The researchers demonstrated that this different behaviour can be traced back to potassium (K) having energetically accessible d-orbitals, which is not the case for lithium and sodium. Under pressure, these orbitals come into play and form different chemical bonds with respect to those observed in lithium and sodium compounds.

Computer-aided design of materials is a new and promising field of materials science. Thanks to modern computing facilities and algorithms, scientists are able to predict the structure and properties of compounds, which will significantly increase the speed and reduce the cost of manufacturing the materials of the future.

More information: G. Saleh et al. Alkali subhalides: high-pressure stability and interplay between metallic and ionic bonds, *Phys. Chem. Chem. Phys.* (2016). [DOI: 10.1039/C5CP06026E](https://doi.org/10.1039/C5CP06026E)

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