

# Using bacteria to accelerate CO<sub>2</sub> capture in oceans

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Berkeley Lab researcher Peter Agbo was awarded a grant for a carbon capture project under the Lab's Carbon Negative Initiative. Credit: Marilyn Sargent/Berkeley Lab

You may be familiar with direct air capture, or DAC, in which carbon dioxide is removed from the atmosphere in an effort to slow the effects

of climate change. Now a scientist at Lawrence Berkeley National Laboratory (Berkeley Lab) has proposed a scheme for direct ocean capture. Removing CO<sub>2</sub> from the oceans will enable them to continue to do their job of absorbing excess CO<sub>2</sub> from the atmosphere.

Experts mostly agree that combating climate change will take more than halting emissions of climate-warming gases. We must also remove the [carbon dioxide](#) and other [greenhouse gases](#) that have already been emitted, to the tune of gigatons of CO<sub>2</sub> removed each year by 2050 in order to achieve net zero emissions. The oceans contain significantly more CO<sub>2</sub> than the atmosphere and have been acting as an important carbon sink for our planet.

Peter Agbo is a Berkeley Lab staff scientist in the Chemical Sciences Division, with a secondary appointment in the Molecular Biophysics and Integrated Bioimaging Division. He was awarded a grant through Berkeley Lab's Carbon Negative Initiative, which is aiming to develop breakthrough negative emissions technologies, for his ocean capture proposal. His co-investigators on this project are Steven Singer at the Joint BioEnergy Institute and Ruchira Chatterjee, a scientist in the Molecular Biophysics and Integrated Bioimaging Division of Berkeley Lab.

## **Q. Can you explain how you envision your technology to work?**

What I'm essentially trying to do is convert CO<sub>2</sub> to limestone, and one way to do this is to use [seawater](#). The reason you can do this is because limestone is composed of magnesium, or what's called magnesium and calcium carbonates. There's a lot of magnesium and calcium naturally resident in seawater. So if you have free CO<sub>2</sub> floating around in seawater, along with that magnesium and calcium, it will naturally form

limestone to a certain extent, but the process is very slow—borderline geologic time scales.

It turns out that the bottleneck in the conversion of CO<sub>2</sub> to these magnesium and calcium carbonates in seawater is a process that is naturally catalyzed by an enzyme called [carbonic anhydrase](#). It's not important to know the enzyme name; it's just important to know that when you add carbonic anhydrase to this seawater mixture, you can basically accelerate the conversion of CO<sub>2</sub> to these limestones under suitable conditions.

And so the idea is to scale this up—drawing CO<sub>2</sub> out of the atmosphere into the ocean and ultimately into some limestone product that you could sequester.

**Q. Fascinating. So you want to turn carbon dioxide into rock using a process that occurs naturally in seawater, but accelerating it. This sounds almost like science fiction. What are the challenges in getting this to work?**

To absorb CO<sub>2</sub> from the air quick enough for the technology to work, you have to solve the problem of how to provide enough of this enzyme that you could deploy this process at a meaningful scale. If we were to simply try to supply the enzyme as a pure product, you couldn't do it in an economically viable way. So the question I'm trying to answer here is, how would you do this? You also have to find ways of stabilizing the pH and mixing in enough air to raise and maintain your CO<sub>2</sub> concentration in water.

The solution that occurred to me was, okay, given that we know carbonic anhydrase is a protein, and proteins are naturally synthesized by

biochemical systems, such as bacteria, which we can manipulate, then we could take bacteria and then engineer them to make carbonic anhydrase for us. And you can just keep growing these bacteria as long as you feed them. One problem, though, is that now you've shifted the cost burden onto supplying enough food to produce enough bacteria to produce enough enzyme.

One way around this issue would be to use bacteria that can grow using energy and nutrients that are readily available in the natural environment. So this pointed towards [photosynthetic bacteria](#). They can use sunlight as their energy source, and they can also use CO<sub>2</sub> as their carbon source to feed on. And certain photosynthetic bacteria can also use the minerals that naturally occur in seawater essentially as vitamins.

**Q. Interesting. So the path to capturing excess CO<sub>2</sub> lays in being able to engineer a microbe?**

Potentially one way, yes. What I've been working on in this project is to develop a genetically modified bacterium that is photosynthetic and is engineered to produce a lot of carbon anhydrase on its surface. Then, if you were to put it in seawater, where you have a lot of magnesium and calcium, and also CO<sub>2</sub> present, you would see a rapid formation of limestone. That's the basic idea.

It's a small project for now, so I decided to focus on getting the engineered organism. Right now, I'm simply trying to develop the primary catalyst system, which are the enzyme-modified bacteria to drive the mineralization. The other non-trivial pieces of this approach—how to appropriately design the reactor to stabilize CO<sub>2</sub> concentrations and pH needed for this scheme to work—are future challenges. But I've been using simulations to inform my approaches to those problems.

It's a fun project because on any given day my co-PIs and I could be doing either physical electrochemistry or gene manipulation in the lab.

## **Q. How would this look once it's scaled up? And how much carbon would it be able to sequester?**

What I have envisioned is, the bacterium would be grown in a plant-scaled bioreactor. You basically flow seawater into this bioreactor while actively mixing in air, and it processes the seawater, converting it to limestone. Ideally, you probably have some type of downstream centrifugation process to extract the solids, which maybe could be driven by the flow of water itself, which then helps to pull out the limestone carbonates before you then eject the depleted seawater. An alternative that could possibly resolve the pH constraints of mineralization would be to implement this instead as a reversible process, where you also use the enzyme to reconvert the carbon you've captured in seawater back to a more concentrated CO<sub>2</sub> stream (carbonic anhydrase behavior is reversible).

What I've calculated for this system, assuming that the protein carbonic anhydrase behaves on the bacterial surface, more or less, the way it does in free solution, would suggest that you would need a plant that has only about a 1-million-liter volume, which is actually quite small. One of those could get you to roughly 1 megaton of CO<sub>2</sub> captured per year. A lot of assumptions are built into that sort of estimate though, and it's likely to change as work advances.

Erecting 1,000 such facilities globally, which is a small number compared to the 14,000 water treatment facilities in the United States alone, would permit the annual, gigaton-scale capture of atmospheric CO<sub>2</sub>.

Provided by Lawrence Berkeley National Laboratory

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