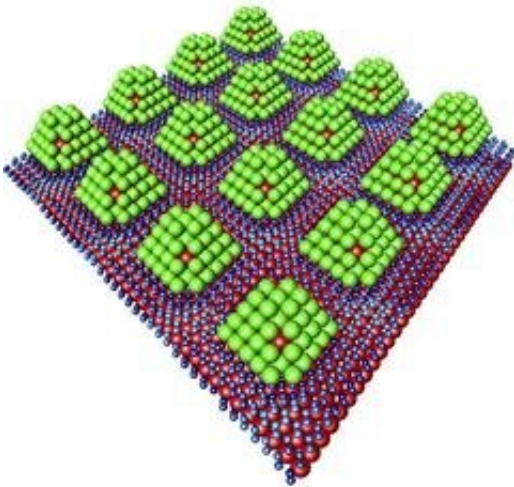


# Team proposes 'nano-chocolates' as a new way to store hydrogen

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The palladium nanoparticles (green) are stabilized by a core of iridium (red). Hydrogen can accumulate on their surface like a kind of chocolate glaze—and can be released again by heating. Credit: DESY, Andreas Stierle

An innovative approach could turn nanoparticles into simple reservoirs for storing hydrogen. The highly volatile gas is considered a promising energy carrier for the future, which could provide climate-friendly fuels for airplanes, ships and trucks, for example, as well as allowing climate-friendly steel and cement production—depending on how the hydrogen gas is generated. However, storing hydrogen is costly: Either the gas has to be kept in pressurized tanks, at up to 700 bar, or it must be liquified, which means cooling it down to minus 253 degrees Celsius. Both

procedures consume additional energy.

A team led by Deutsches Elektronen-Synchrotron (DESY)'s Andreas Stierle has laid the foundations for an alternative method: Storing [hydrogen](#) in tiny nanoparticles made of the precious metal [palladium](#), just 1.2 nanometers in diameter. The fact that palladium can absorb hydrogen like a sponge has been known for some time. "However, until now getting the hydrogen out of the material again has posed a problem," Stierle explains. "That's why we are trying palladium particles that are only about one nanometer across." A nanometer is a millionth of a millimeter.

To ensure that the [tiny particles](#) are sufficiently sturdy, they are stabilized by a core made of the rare precious metal iridium. In addition, they are attached to a graphene support, an extremely thin layer of carbon. "We are able to attach the palladium particles to the graphene at intervals of just two and a half nanometers," reports Stierle, who is the head of the DESY NanoLab. "This results in a regular, periodic structure." The team, which also includes researchers from the Universities of Cologne and Hamburg, published its findings in the American Chemical Society (ACS) journal *ACS Nano*.

DESY's X-ray source PETRA III was used to observe what happens when the palladium particles come into contact with hydrogen: Essentially, the hydrogen sticks to the nanoparticles' surfaces, with hardly any of it penetrating inside. The nanoparticles can be pictured as resembling chocolates: An iridium nut at the center, enveloped in a layer of palladium, rather than marzipan, and chocolate-coated on the outside by the hydrogen. All it takes to recover the stored hydrogen is for a small amount of heat to be added; the hydrogen is rapidly released from the surface of the particles, because the gas molecules don't have to push their way out from inside the cluster.

"Next, we want to find out what storage densities can be achieved using this new method," says Stierle. However, some challenges still need to be overcome before proceeding to practical applications. For example, other forms of carbon structures might be a more suitable carrier than graphene—the experts are considering using carbon sponges, containing tiny pores. Substantial amounts of the palladium nanoparticles should fit inside these.

The latest issue of DESY's research magazine, *femto*, looks into this and other innovative concepts for the hydrogen economy and a sustainable energy supply. The magazine explains how fundamental research can contribute to innovations for the energy transition. This is not only about using hydrogen as an energy carrier, but also about sustainable solar cells and novel forms of energy generation, as well as achieving greater energy efficiencies in research itself, when operating large particle accelerators, for instance.

**More information:** Dirk Franz et al, Hydrogen Solubility and Atomic Structure of Graphene Supported Pd Nanoclusters, *ACS Nano* (2021).  
[DOI: 10.1021/acsnano.1c01997](https://doi.org/10.1021/acsnano.1c01997)

Provided by Deutsches Elektronen-Synchrotron

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