

# Recent advances in the optimization of dispersed platinum catalysts

June 10 2021

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Native Platinum nugget, locality Kondyor mine, Khabarovsk Krai, Russia.  
Credit: Wikipedia.

Highly dispersed platinum catalysts provide new possibilities for industrial processes, such as the flameless combustion of methane, propane, or carbon monoxide, which has fewer emissions and is more resource efficient and consistent than conventional combustion. In the journal *Angewandte Chemie*, a team of researchers reports on which platinum species are active in high-temperature oxidations and what changes they can undergo in the course of the process—important prerequisites for the optimization of catalysts.

Individual metal [atoms](#) and clusters consisting of only a few metal atoms have interesting catalytic properties determined by the exact nature of the active metal species. Usually, these are highly dispersed and

deposited on a support such as zeolite, which is a porous silicate framework structure that also plays a role in the characteristics of a [catalyst](#). Even the smallest change in the active centers can drastically reduce the efficiency of a catalyst. For example, noble metals like [platinum](#) tend to become permanently deactivated through sintering under harsh conditions.

Which specific platinum species play a role in high-temperature oxidations is hard to determine, however, because a significant number of such species cannot readily be obtained without the involvement of their support in the catalysis. A team led by Pedro Serna (ExxonMobil Research and Engineering Co., New Jersey, U.S.), as well as Manuel Moliner and Avelino Corma (Universitat Politècnica de València, Spain) investigated the behavior of individual platinum atoms and small platinum clusters on special CHA zeolites, which are non-reducible supports that can stabilize these species very well.

Their first experiment was an investigation of splitting  $O_2$  using two different types of isotopically pure oxygen molecules,  $^{16}O_2$  and  $^{18}O_2$ . The more active the catalyst, the more mixed  $^{16}O(^{18}O)$  molecules are formed upon recombination of the dissociated atoms. It was shown that platinum clusters of under one nanometer are significantly more active than individual atoms or larger clusters. However, at moderate temperatures (200 °C) the tiny clusters fall apart over time into individual platinum atoms and the [catalytic activity](#) for splitting oxygen ends.

In contrast, the team found that for the [oxidation](#) of alkanes, such as methane, at higher temperatures, the catalytic combustion was carried out by individual platinum atoms. These are formed in situ in the oxygen stream from the initial clusters, as was shown by X-ray absorption spectroscopy and by electron microscopy. The critical step in these oxidations is not the splitting of  $O_2$  but the breaking of C-H bonds,

which is less sensitive to changes in the active-site structure.

For the oxidation of CO, the catalysis is dominated by platinum clusters. Individual platinum atoms cannot be stabilized in the CO stream, and thus, play no role. In comparison with supports made of aluminum oxide, the CHA zeolite provided higher activity and greater stability of the platinum clusters in the presence of CO.

The high stability of individual platinum atoms for methane combustion and of small platinum clusters for CO oxidation, which is retained after regeneration or treatment with hot steam, opens new possibilities for systems made of platinum and silicate zeolites as efficient and robust heterogeneous catalysts for a variety of high-temperature oxidation scenarios.

**More information:** Pedro Serna et al, Single-Site vs. Cluster Catalysis in High Temperature Oxidations, *Angewandte Chemie International Edition* (2021). [DOI: 10.1002/anie.202102339](https://doi.org/10.1002/anie.202102339)

Provided by Wiley

Citation: Recent advances in the optimization of dispersed platinum catalysts (2021, June 10) retrieved 20 April 2024 from <https://phys.org/news/2021-06-advances-optimization-dispersed-platinum-catalysts.html>

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