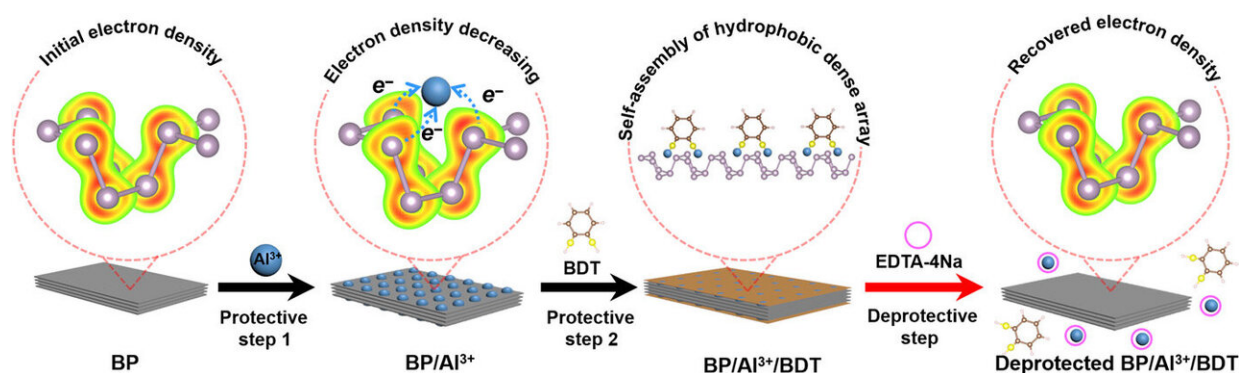


# Regulating the reactivity of black phosphorous through protective chemistry

November 24 2020, by Thamarasee Jeewandara



Schematic illustration of regulating the reactivity of BP via protective chemistry. Protective step 1: Binding  $\text{Al}^{3+}$  ions with lone pair electrons on the surface of P atoms decreases surface electron density of BP, leading to a reduced chemical reactivity of BP. Protective step 2: Self-assembly of the hydrophobic dense array on the BP surface isolates BP from surrounding oxygen/water. Deprotective step: Removal of  $\text{Al}^{3+}$  ions and hydrophobic dense array on the BP surface by a chelating agent. The treatment recovers the electron density of BP, restoring the original reactivity of the deprotected BP. BDT, 1,2-benzenedithiol; EDTA-4Na, EDTA-tetrasodium. Credit: Science Advances, doi: 10.1126/sciadv.abb4359.

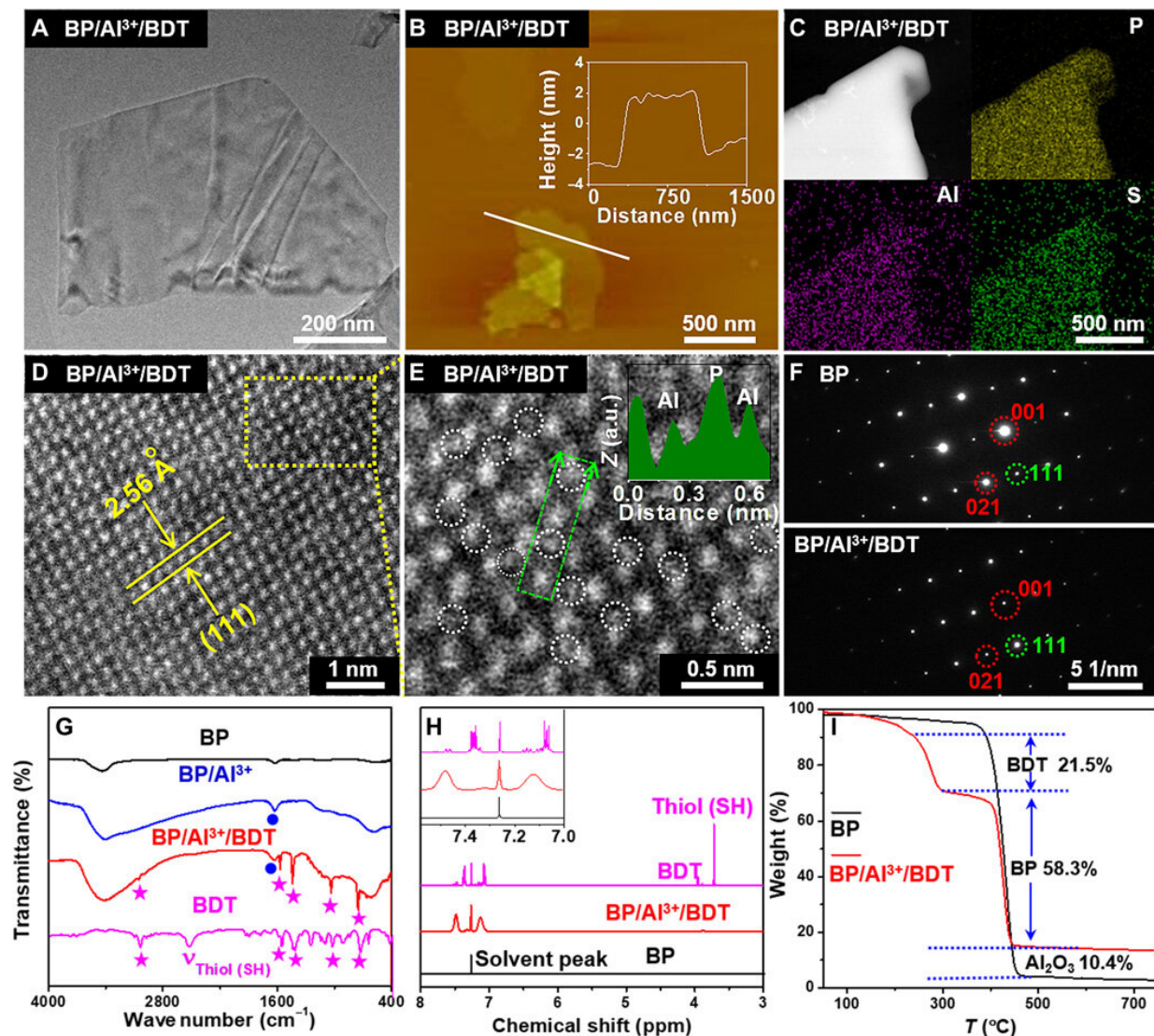
Chemists can rationally regulate the reactivity of molecules and functional groups in both industrial and laboratory-based synthetic organic chemistry processes. The concept can be applied to inorganic nanomaterials including two-dimensional (2-D) [black phosphorus](#) (BP) nanosheets. For example, scientists can "shut down" the high reactivity

of few layer or monolayer black phosphorus, when the compound is not in use, to resume its activity upon application. In a new report now published on *Science Advances*, Xiao Liu and a team of scientists in physics, biomaterials, chemical engineering and biological chemistry in China developed a protective chemistry-based method to regulate the reactivity of black phosphorus.

To initiate the protective step, they bound [aluminum cations](#) ( $\text{Al}^{3+}$ ) with lone pairs of electrons from black phosphorous (BP) and decreased the [electron density](#) on the black [phosphorus](#) surface. They completed the process with an oxygen/water-resistant layer via the self-assembly of hydrophobic (water-hating) [1,2-benzenedithiol](#) (BDT) on the black phosphorus/aluminum ( $\text{BP}/\text{Al}^{3+}$ ) conjugate. The protective step produced a stable compound with low reactivity. By using a [chelator treatment](#) process, Liu et al. subsequently achieved deprotection of the  $\text{BP}/\text{Al}^{3+}/\text{BDT}$  complex to remove aluminum cations and BDT from the black phosphorus surface. In this way, they recovered the electron density of black phosphorus using the de-protective process to restore reactivity to the compound.

## Tuning the properties of nanomaterials

Liu et al. rationally developed a protective chemistry-based method to control the reactivity of black phosphorus during this work. In nanoscience, researchers can precisely [tune the properties of nanomaterials](#) to obtain desired characteristics. Regulating the reactivity of nanomaterials is critical for multistep programmable applications. Some nanomaterials can be protected to reduce their reactivity under specific conditions and restore activity after successful deprotection. Researchers have therefore [proposed a range of selective and efficient](#) strategies to regulate the reactivity of functional groups in organic chemistry.



Characterization of BP/Al<sup>3+</sup>/BDT. (A) TEM image. (B) AFM (height profile along the white line) image. (C) STEM–energy-dispersive x-ray spectroscopy (EDX) elemental mapping images. (D) High-angle annular dark-field (HAADF) image. (E) Magnified HAADF image taken from the selected area in (D). a.u.: arbitrary units. (F) Selected-area electron diffraction (SAED) pattern of BP and BP/Al<sup>3+</sup>/BDT. (G) FTIR spectra of BP, BP/Al<sup>3+</sup>, BP/Al<sup>3+</sup>/BDT, and BDT. (H) <sup>1</sup>H NMR spectra of BP, BP/Al<sup>3+</sup>/BDT, and BDT. (I) Thermogravimetric curves of BP and BP/Al<sup>3+</sup>/BDT. ppm, parts per million. Credit: Science Advances, doi: 10.1126/sciadv.abb4359.

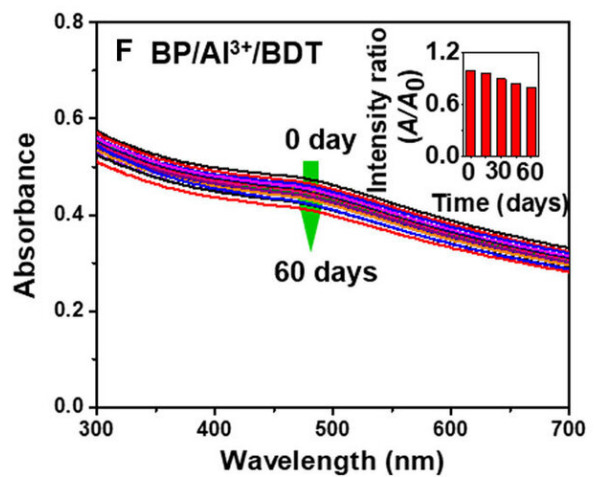
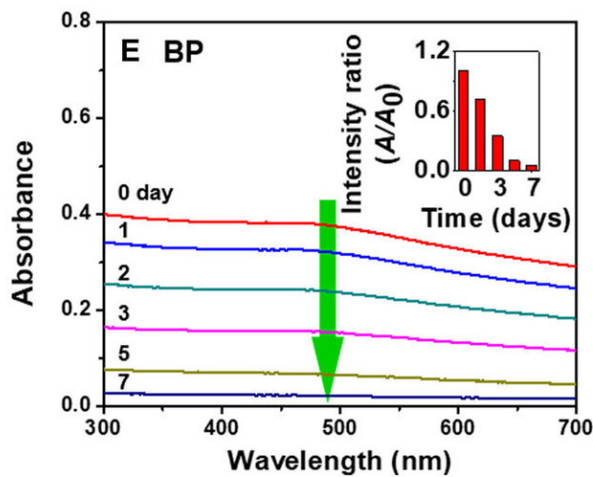
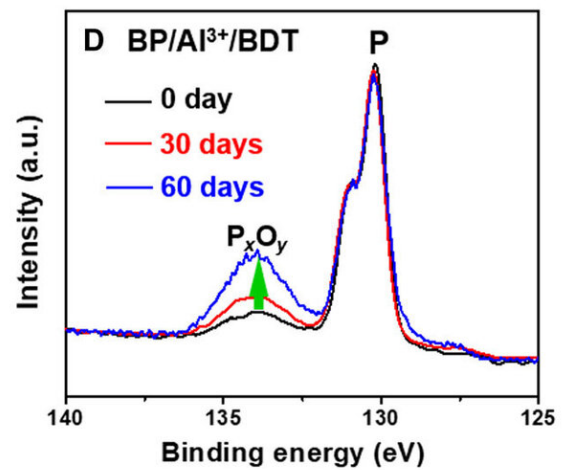
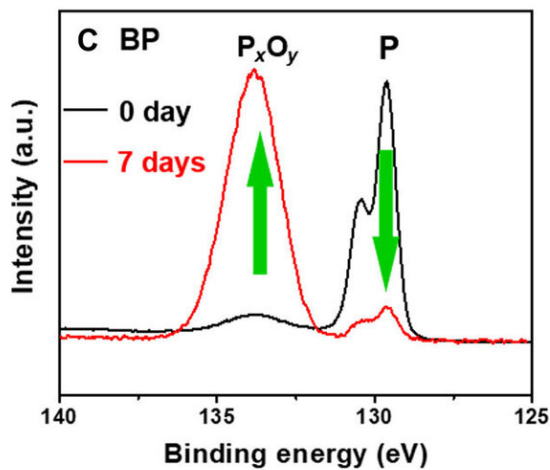
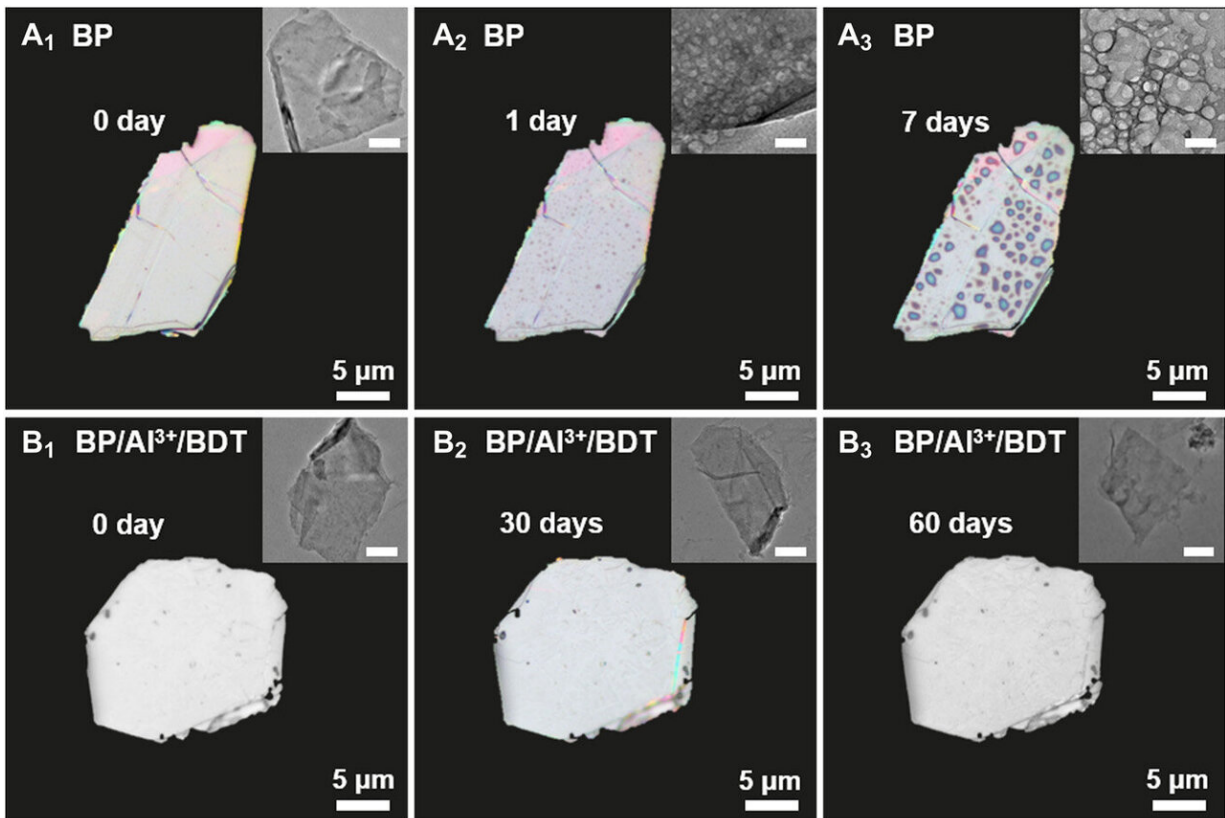
However, the existing [well-established organic protective-deprotective processes](#) cannot be applied to inorganic nanomaterials due to a lack of similar functional groups. As a result, an efficient and simple approach to control the reactivity of inorganic materials remains to be developed. To accomplish this, the team began binding black phosphorus with aluminum cations ( $\text{Al}^{3+}$ ) to decrease the surface electron density and effectively decrease its reactivity. The protective process produced an arrangement of the hydrophobic (water-hating) 1,2-benzendithiol (BDT) molecule in the presence of black phosphorus and  $\text{Al}^{3+}$  cations to offer an ultra-stable complex (BP/  $\text{Al}^{3+}$  /BDT). The compound was stable under ambient conditions for up to two months without alterations. The team can deprotect the unstable complex through chelator treatment.

### **Synthesis and characterization of the BP/ $\text{Al}^{3+}$ /BDT compound**

Liu et al. synthesized and characterized (tested) the bulk black phosphorous (BP) by following a [previously developed method](#). The team first obtained BP nanosheets by sonicating the powdered form of the molecule, then using scanning [electron microscopy](#) and [transmission electron microscopy](#) (TEM), they observed the size of BP. High resolution TEM too provided insight to the nanosheet structure and [atomic force microscopy](#) revealed the thickness of BP with four to six individual layers of [phosphorene layers](#). Using X-ray diffraction and Raman spectra, the team determined the crystal characteristics of BP to be similar to its bulk form. At the protective step of cationic binding to the BP surface, Liu et al. mixed BP with aluminum chloride ( $\text{AlCl}_3$ ) in an ethanol solution. They then characterized the successful attachment of  $\text{Al}^{3+}$  cations to the BP surface to strengthen the protection of BP by employing [X-ray photoelectron spectroscopy](#). The team observed the nano-morphology of the compound and studied the surface conformation using [scanning transmission electron microscopy](#) (STEM) and further verified its structure with [Fourier transform infrared](#) (FTIR)

spectroscopy and [proton nuclear magnetic resonance](#) ( $^1\text{H}$ NMR).

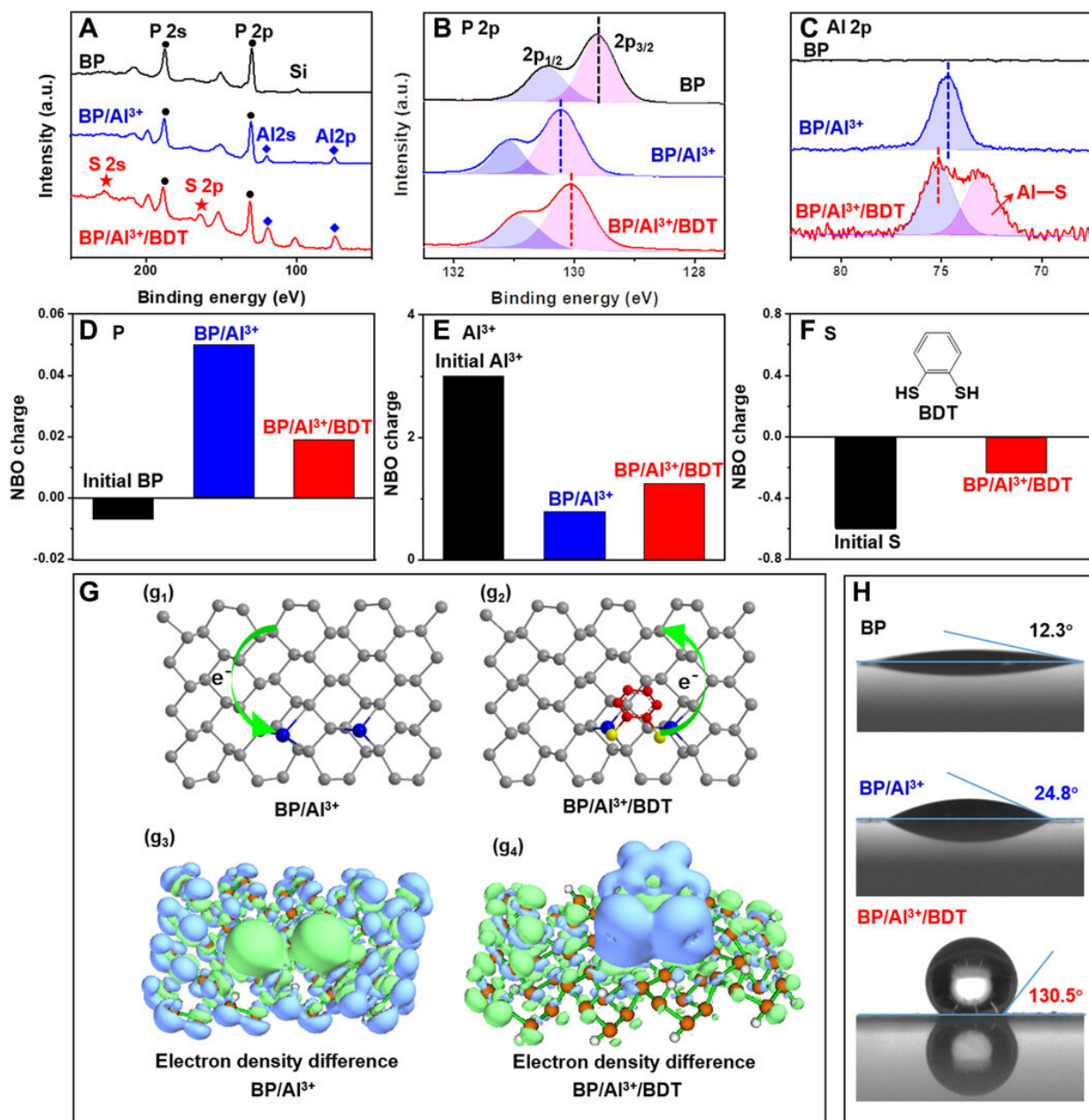




Characterization of degraded BP and BP/Al<sup>3+</sup>/BDT under ambient conditions. Polarizing microscope images of (A) bulk BP (0, 1, and 7 days) and (B) bulk BP/Al<sup>3+</sup>/BDT (0, 30, and 60 days). Insets: Corresponding TEM images. Scale bars, 200 nm. (C and D) HR-XPS spectra of P 2p peaks for BP and BP/Al<sup>3+</sup>/BDT with ambient exposure for various durations. (E and F) UV-vis spectra of BP and BP/Al<sup>3+</sup>/BDT dispersed in water for various durations. Insets: variation of the UV-vis absorption ratios at 470 nm (A/A<sub>0</sub>) of BP (A<sub>0</sub>: original value). Credit: Science Advances, doi: 10.1126/sciadv.abb4359

### **Regulating the reactivity of the BP/Al<sup>3+</sup>/BDT compound**

Liu et al. investigated the reactivity of the BP/Al<sup>3+</sup>/BDT complex by employing a variety of optical techniques in materials science to understand its structure. The corresponding images showed the evolution of the crystal structure and long-term stability under ambient conditions. The stability of the BP/Al<sup>3+</sup>/BDT complex was superior to black phosphorus alone and the results highlighted the practicality of embedding black phosphorus in the compound. The team credited the reduced reactivity of the complex structure to two factors; first, the aluminum cation binding the black phosphorus surface rendered lower chemical reactivity. Second, the self-assembled hydrophobic dense array on the surface of black phosphorus isolated the molecule from oxygen and water to prevent further degradation, improving the stability of the compound.



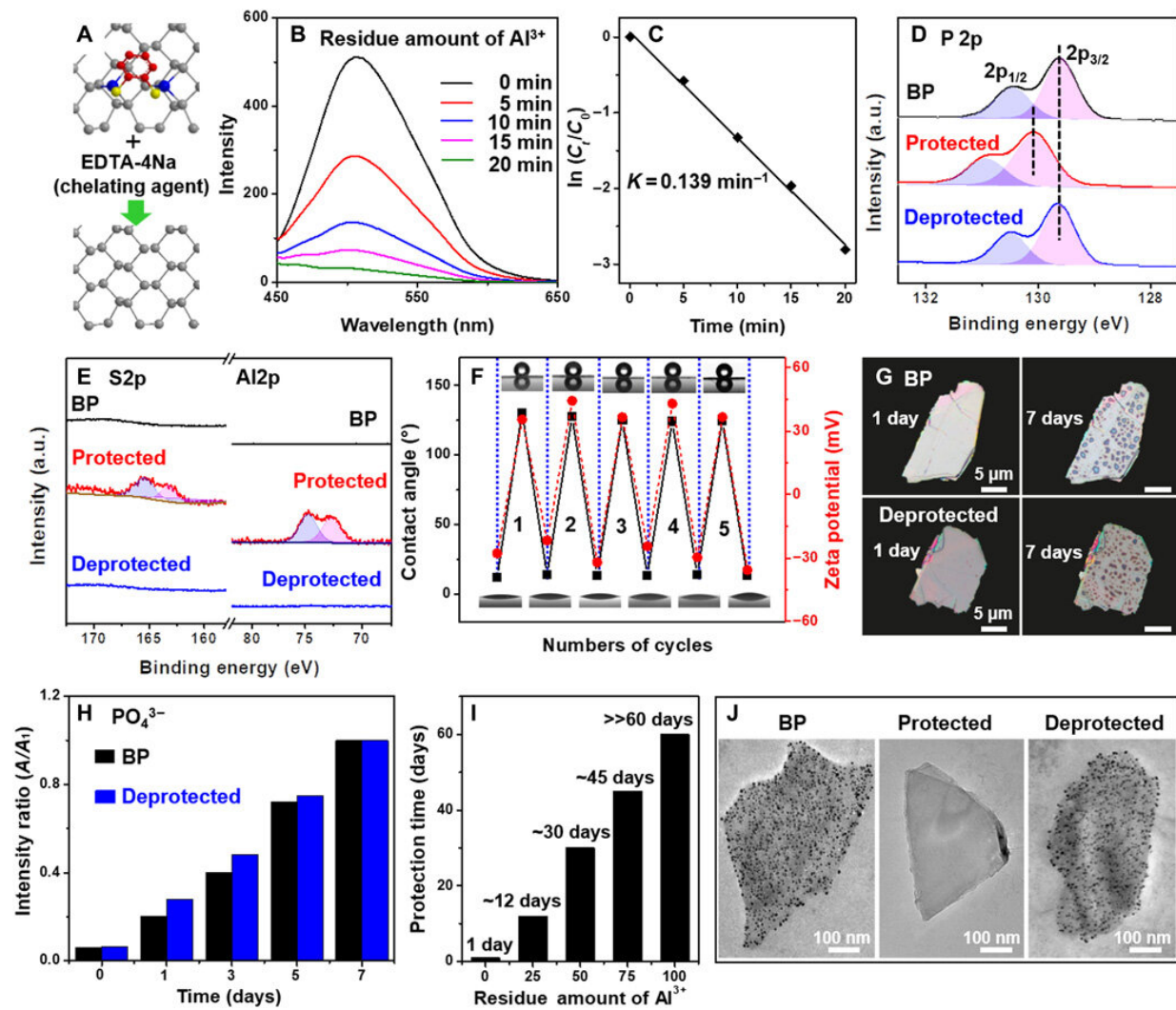
Mechanism of reactivity decrease in protected BP. (A) Full XPS spectra of BP, BP/Al<sup>3+</sup>, and BP/Al<sup>3+</sup>/BDT. (B and C) HR-XPS spectra of P 2p and Al 2p. (D to F) Calculated NBO charge of P atom, Al<sup>3+</sup> ion, and S atom. Structure model of (G<sub>1</sub>) BP/Al<sup>3+</sup> and (G<sub>2</sub>) BP/Al<sup>3+</sup>/BDT. Computational mapping of electron density difference in (G<sub>3</sub>) BP/Al<sup>3+</sup> and (G<sub>4</sub>) BP/Al<sup>3+</sup>/BDT. Green regions indicate increased electron density, and blue regions indicate decreased electron density. Contours are shown at the 0.0001 a.u. level. (H) Water contact angles of BP, BP/Al<sup>3+</sup>, and BP/Al<sup>3+</sup>/BDT. Credit: Science Advances, doi:



10.1126/sciadv.abb4359.

## Deprotecting the BP/Al<sup>3+</sup>/BDT compound

The research team deprotected the ultra-stable complex by removing the aluminum cation from the surface of the compound. They accomplished this step with a conventional metal ion chelator—[edetate sodium](#) (EDTA-4Na). During the process, Liu et al. also removed the hydrophobic BDT molecules together with aluminum cations, to obtain the resulting hydrophilic (water-loving) surface with [negative zeta potential](#), similar in form to the original black phosphorus molecule. The protective-deprotective regulation process allowed the scientists to reversibly control the reactivity of black phosphorus. The results suggest the ability to efficiently regulate the reactivity of the molecule in practice.



Deprotection of BP/ $\text{Al}^{3+}$ /BDT. (A) Schematic illustration of  $\text{Al}^{3+}$  ion and BDT removal by EDTA-4Na. (B) Photoluminescence (PL) emission spectra of  $\text{Al}^{3+}$  residue on BP/ $\text{Al}^{3+}$ /BDT after EDTA-4Na treatment. (C) Plot of  $\ln(C_t/C_0)$  as a function of EDTA-4Na treatment time. (D and E) HR-XPS spectra of P 2p, Al 2p, and S 2p for BP, BP/ $\text{Al}^{3+}$ /BDT, and deprotected BP/ $\text{Al}^{3+}$ /BDT. (F) Plots of water contact angles and zeta potentials of BP as measured at each protective-deprotective cycle. (G) Polarizing microscope images of bulk BP (0 and 7 days) and bulk deprotected BP/ $\text{Al}^{3+}$ /BDT (0 and 7 days). (H) Variation of  $\text{PO}_4^{3-}$  concentration in solutions of BP and deprotected BP/ $\text{Al}^{3+}$ /BDT with varying ambient exposure durations. (I) Stability of deprotected BP/ $\text{Al}^{3+}$ /BDT with a varying residual amount of  $\text{Al}^{3+}$  ion on the BP surface. (J) TEM images of BP, BP/ $\text{Al}^{3+}$ /BDT, and deprotected BP/ $\text{Al}^{3+}$ /BDT after H $\text{AuCl}_4$  (aqueous solution)

treatment. Credit: Science Advances, doi: 10.1126/sciadv.abb4359.

## Outlook

In this way, Xiao Liu, and colleagues developed a new protective chemistry-based approach to rationally regulate the reactivity of black phosphorus (BP). During the protective step, they bound aluminum cations ( $\text{Al}^{3+}$ ) to the lone pair of electrons on BP to decrease its surface electron density. This protective step enhanced the functionalization of the hydrophobic (water-hating) 1,2-benzendithiol (BDT), on the BP/ $\text{Al}^{3+}$  conjugate surface to form a dense hydrophobic layer, which greatly reduced the reactivity of black phosphorus. The team then used a chelator to remove the aluminum cations from [black phosphorus](#) to revert the molecule back to its original high electron density, hydrophilic (water-loving) [surface](#). The deprotected BP showed high reactivity in its original state. The method provided a tunable approach to manipulate the reactivity of BP, which is hard to achieve through conventional functionalization. This protective strategy can be explored to regulate the [reactivity](#) of nanomaterials to create futuristic programmable nanostructures for applications in materials science.

**More information:** Liu X. et al. Regulating the reactivity of black phosphorus via protective chemistry, *Science Advances*, 10.1126/sciadv.abb4359

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