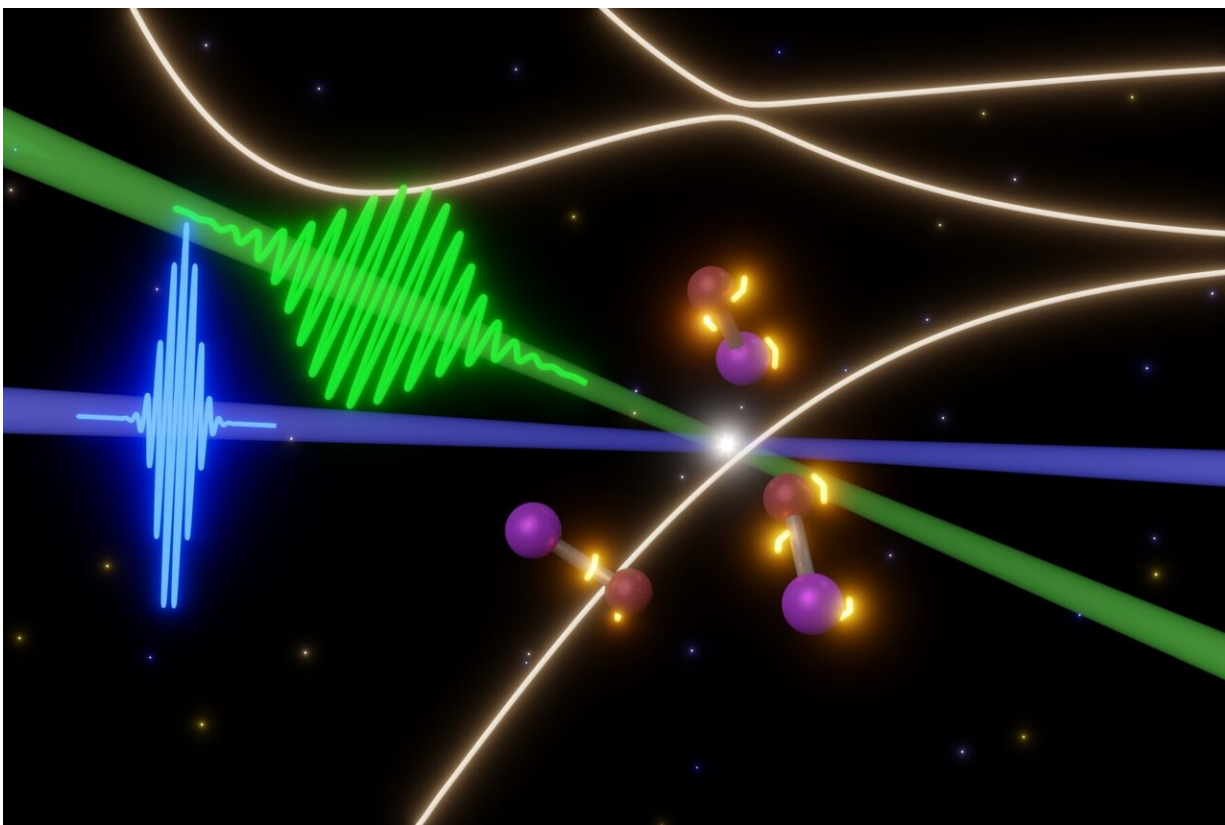


# What happens when you explode a chemical bond?

July 11 2019, by Robert Sanders



UC Berkeley scientists are probing the fleeting steps in rapid photochemical reactions with some of the shortest laser pulses possible today. In this case, a femtosecond pulse of visible light (green) triggers the breakup of iodine monobromide molecules (center), while attosecond XUV laser pulses (blue) take snapshots of the molecules. This allows them to make a movie of the evolution of electronic states (yellow lights around molecules) before the molecules blow apart. Credit: Yuki Kobayashi, UC Berkeley

On bright summer days, the sunlight all around us is breaking bad by breaking bonds. Chemical bonds.

Ultraviolet [light](#) shatters the links between atoms in the DNA of our skin cells, potentially causing cancer. UV light also breaks oxygen bonds, eventually creating ozone, and cleaves hydrogen off other [molecules](#) to leave behind free radicals that can damage tissue.

University of California, Berkeley, chemists using some of the shortest laser pulses available—one quintillionth of a second—have now been able to resolve the step-by-step process leading to the exploding of a [chemical bond](#), essentially making a movie of the event. They can follow electrons indecisively bouncing around in various states in the molecule before the bond breaks, and the atoms go their separate ways.

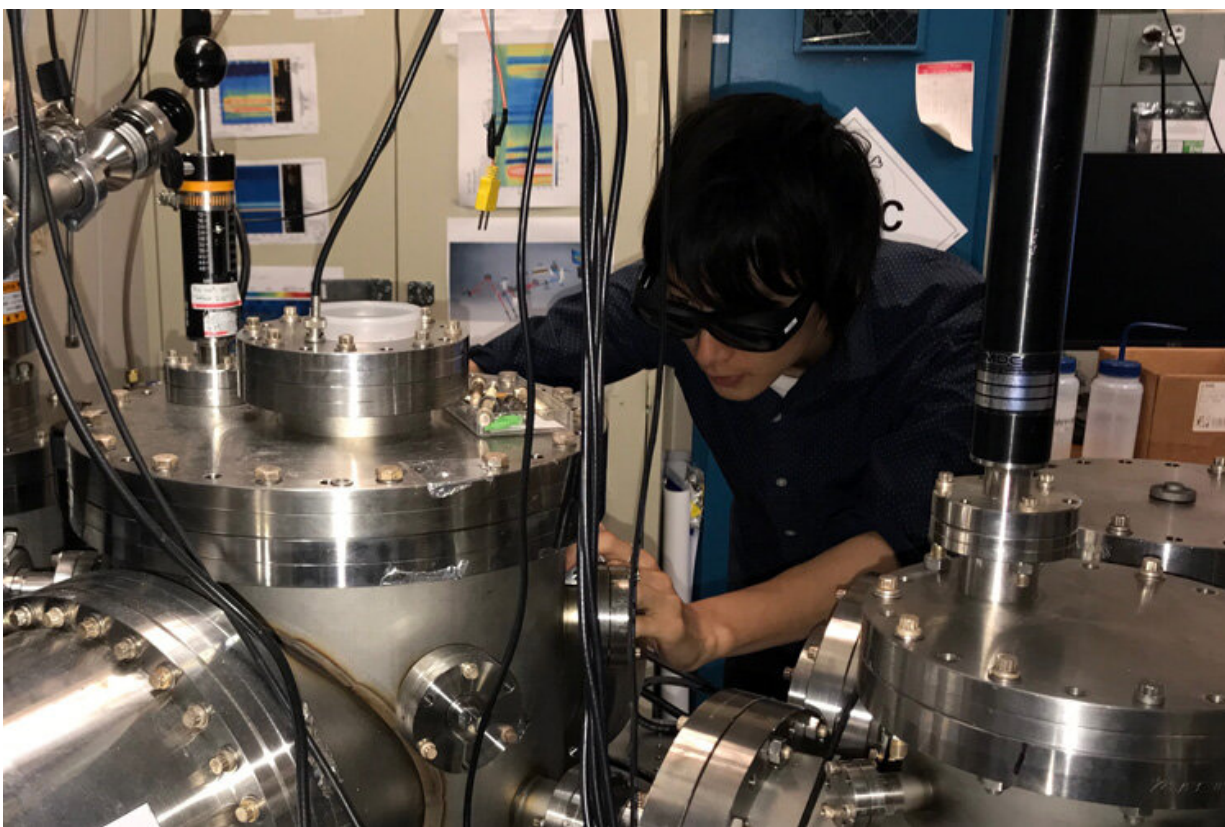
The technique, reported last week in the journal *Science*, will help chemists understand and potentially manipulate [chemical reactions](#) stimulated by light, so-called photochemical reactions. Chemists and biologists, in particular, are interested in understanding how [large molecules](#) manage to absorb [light energy](#) without breaking any bonds, as happens when molecules in the eye absorb light, giving us vision, or molecules in plants absorb light for photosynthesis.

"Think about a molecule, rhodopsin, in the eye," said first author Yuki Kobayashi, a UC Berkeley doctoral student. "When light hits the retina, rhodopsin absorbs the [visible light](#), and we can see things because rhodopsin's bond's conformation changes."

In fact, when the light energy is absorbed, a bond in rhodopsin twists, instead of breaks, triggering other reactions that result in the perception of light. The technique Kobayashi and his UC Berkeley colleagues, professors Stephen Leone and Daniel Neumark, developed could be used to study in detail how this light absorption leads to twisting after the

molecule passes through an [excited state](#) called an avoided crossing or conical intersection.

To prevent the breaking of a bond in DNA, "you want to redirect the energy from dissociation to just being vibrationally hot. For rhodopsin, you want to redirect the energy from vibrating to a cis-trans isomerization, a twist," Kobayashi said. "These redirections of chemical reactions are happening ubiquitously around us, but we have not seen the actual moment of them before."



Yuki Kobayashi makes adjustments to an experiment in the basement of Hildebrand Hall. He and researchers in the labs of Stephen Leone and Daniel Neumark target molecules in a vacuum chamber with both femtosecond and attosecond lasers to find out how molecules absorb light. Credit: UC Berkeley photo courtesy of Robert Sanders

## Fast laser pulses

Attosecond lasers—an attosecond is a billionth of a billionth of a second—have been around for about a decade and are used by scientists to probe very fast reactions. Since most chemical reactions occur rapidly, these fast-pulse lasers are key to "seeing" how the electrons that form the chemical bond behave when the bond breaks and/or reforms.

Leone, a professor of chemistry and of physics, is an experimentalist who also uses theoretical tools and is a pioneer in using attosecond lasers to probe chemical reactions. He has six of these X-ray and extreme ultraviolet (collectively, XUV) lasers in his UC Berkeley laboratory.

Working with one of the simplest of molecules, iodine monobromide (IBr)—which is one iodine atom linked to one bromine atom—the UC Berkeley team hit the molecules with an 8 femtosecond pulse of visible light to excite one of their outermost electrons, then probed them with attosecond laser pulses.

Pulsing the attosecond XUV laser at timed intervals of 1.5 femtosecond (a femtosecond is 1,000 attoseconds), much like using a strobe light, the researchers could detect the steps leading to the breakup of the molecules. The high-energy XUV laser was able to explore the excited energy states relative to the molecule's inner electrons, which normally do not participate in chemical reactions.

"You are kind of making a movie of the pathways of the electron when it approaches the crossing and the probability of it going along one path or along another," Leone said. "These tools we are developing allow you to look at solids, gases and liquids, but you need the shorter time scales (provided by an attosecond [laser](#)). Otherwise, you only see the beginning

and the end, and you don't know what else happened in between."

The experiment showed clearly that the outer electrons of IBr, once excited, suddenly see a variety of states or places they could be and explore many of them before deciding which path to take. In this simple molecule, however, all paths lead to the electron settling on either iodine or bromine and the two atoms flying apart.

In larger molecules, which the team hopes soon to explore, excited electrons would have more choices, some where the energy goes into a twist, like with rhodopsin, or into molecular vibration without the molecules breaking apart.

"In biology, it turns out that evolution has selected things that are extremely effective at absorbing the energy and not breaking a bond," Leone said. "When something goes wrong in your chemistry is when you see diseases cropping up."

**More information:** Yuki Kobayashi et al, Direct mapping of curve-crossing dynamics in IBr by attosecond transient absorption spectroscopy, *Science* (2019). [DOI: 10.1126/science.aax0076](https://doi.org/10.1126/science.aax0076)

Provided by University of California - Berkeley

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