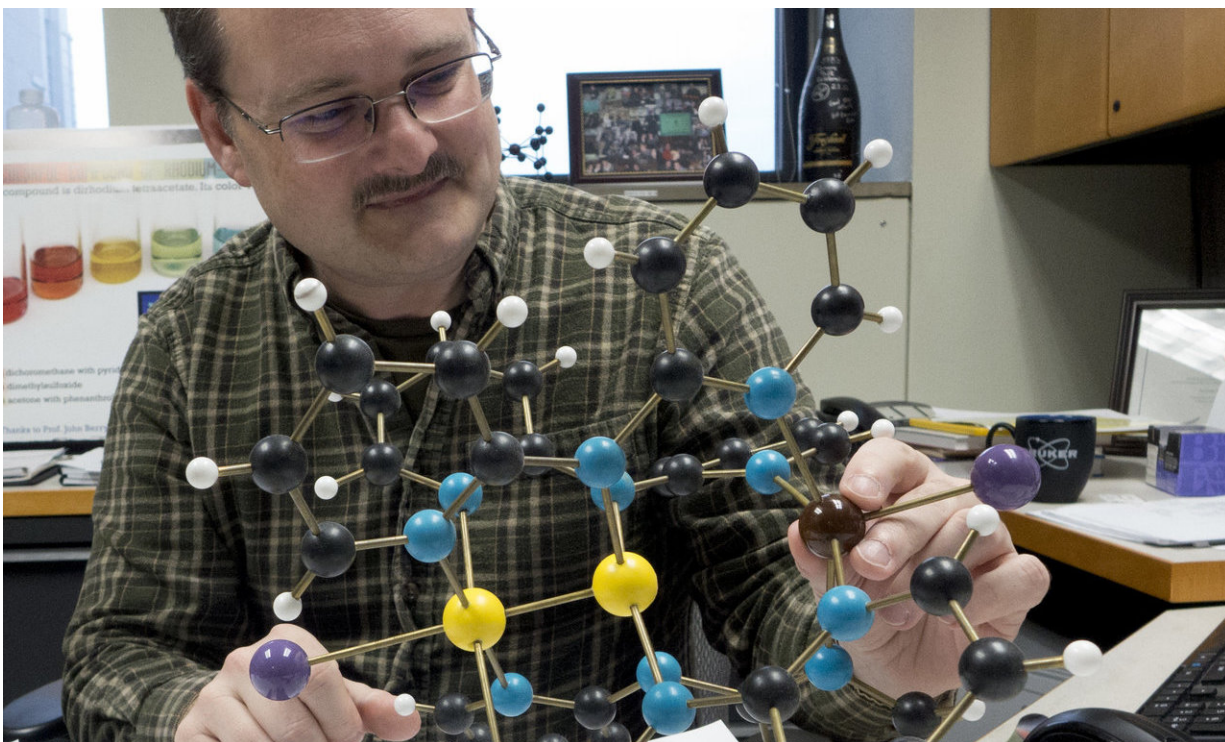


Molecular magnetism packs power with 'messenger electron'

November 13 2017, by David Tenenbaum



Atoms on adjacent molecules like this could be linked to form a long, magnetic chain, creating a new type of magnetic structure, says John Berry, a professor of chemistry at the University of Wisconsin-Madison. Credit: David Tenenbaum, UW-Madison

Electrons can be a persuasive bunch, or at least, a talkative bunch, according to new work from John Berry's lab at the University of

Wisconsin-Madison.

The spins of unpaired electrons are the root of permanent magnetism, and after 10 years of design and re-design, Berry's lab has made a molecule that gains magnetic [strength](#) through an unusual way of controlling those spins.

Berry says the new structure that graduate student Jill Chipman created could lead to a breakthrough in [quantum computing](#), an approach with such great potential that it could undermine today's silicon-based supercomputers much as the telephone did the telegraph: A great leap forward that begins a slide into irrelevance.

The presence and activity, or "spin," of unpaired electrons sets the strength of a [permanent magnet](#), so [molecules](#) with a high degree of spin are a desirable target for chemists. The unusually large spin in the new magnetic molecule, Berry explains, results from a "messenger electron" that shuttles between an unpaired electron at each end of the rod-shaped molecule and persuades all three of them to adopt the same spin.

That agreement of spin, "orthogonality" in the jargon, adds strength to a permanent magnet.

Berry, a UW-Madison professor of chemistry, notes that in other materials, a traveling electron tends to oppose the spins of magnetic centers, reducing the magnetic strength. In Chipman's new creation, however, the messenger electron is focused on harmony: like a traveling social worker, it causes the two remote unpaired electrons to take the same spin, adding strength and/or durability.

The new molecule, described in *Chemistry - A European Journal*, contains carbon, nickel, chlorine, nitrogen, and molybdenum, but lacks the costly rare earth elements that have bedeviled efforts to

commercialize super-strong new magnets. Its structure suggests that the molecule could be formed into a polymer - a repeating chain of units like those found in plastics - raising the possibility of cheaper, stronger magnets.

"We tried to remove electrons from this molecule 10 years ago so it had an unpaired electron at each end, but did not get far," Berry says. "We since learned that this made a chemical that is really temperature-sensitive, so Jill had to develop a low-temperature process that relies on dry ice to cool it to -78 degrees C."

The "traveling social worker" electron establishes "a design principle that could be used to create many new magnetic molecules that behave as little bar magnets," Berry says.

The discovery was also enabled by the arrival last summer an instrument called a SQUID magnetometer (Superconducting QUantum Interference Device) that can measure [magnetism](#) with great accuracy down to below 2 degrees above absolute zero.

Much of the focus of magnet innovation concerns greater strength, Berry says, "but there are all sorts of things people look for. We need both permanent magnets and those with ephemeral magnetization for different technical reasons. Magnets are widespread in ultra-cold refrigeration, motors, computer hard drives and electronic circuits."

By going the next step, and miniaturizing magnets to a single molecule, that could enable quantum computing, Berry says. Quantum computing could be especially beneficial to chemists, who confront staggering complexity in trying to model the chemical reactions that are their bread and butter.

Provided by University of Wisconsin-Madison

Citation: Molecular magnetism packs power with 'messenger electron' (2017, November 13)
retrieved 10 April 2024 from

<https://phys.org/news/2017-11-molecular-magnetism-power-messenger-electron.html>

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