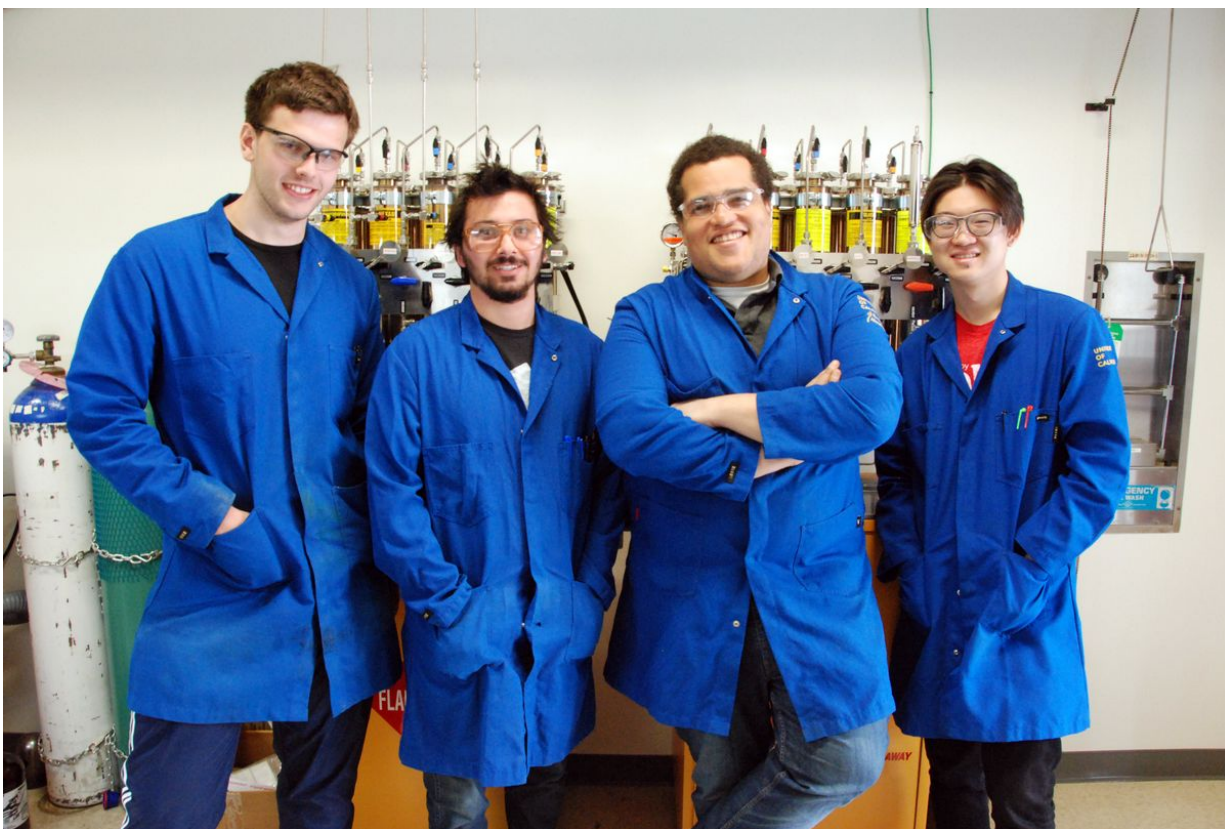


New chemical reaction could eventually yield new fuels and medications

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The researchers from left: Stasik Popov, Alex Bagdasarian, Hosea Nelson and Brian Shao. Credit: Penny Jennings

When scientists develop the chemical formulas for new products such as fuels and medications, they often must first create molecules that haven't

previously existed.

A basic step toward creating new [molecules](#) is selectively breaking and re-forming the chemical bonds that connect the atoms that make them up. One of the chief challenges is that the [bond](#) between carbon and hydrogen atoms—the [building blocks](#) of many molecules—is exceptionally strong, so chemists often have to resort to using rare and expensive chemicals like iridium to convert it into other, more useful types of [chemical bonds](#). Scientists refer to this process as "functionalizing" the bonds.

Now, a team of UCLA chemists has developed a new technique for breaking [carbon-hydrogen bonds](#) and making [carbon-carbon bonds](#). The approach uses catalysts made of two abundant and inexpensive elements, silicon and boron. Their research was published in *Science*.

Hosea Nelson, a UCLA assistant professor of chemistry and biochemistry and senior author of the study, said the energy industry has been interested in taking very simple hydrocarbon molecules like methane and turning them into new fuels.

"This new method will enable scientists to incorporate methane into bigger molecules," he said.

Another potential application would be converting methane, one of the primary components of natural gas, into something that's denser and easier to contain after it has been drilled from the Earth. The current process is complicated because methane, a light gas, tends to escape into the atmosphere.

Nelson collaborated on the study with UCLA graduate students Brian Shao, Alex Bagdasarian and Stasik Popov.

The researchers used their new technique to create a compound similar to a phenyl cation, a chemical substance that has been studied theoretically but rarely investigated in actual laboratory experiments. They then used the compound to slice through carbon-hydrogen bonds in methane and benzene, which allowed them to insert other atoms and form carbon-carbon bonds, which are the basic building blocks of molecules that make up living organisms, as well as fuels and pharmaceuticals.

Besides demonstrating that phenyl cation-like compounds exist, the [new technique](#) allows complex molecules to be assembled in far fewer reaction steps than was previously possible, which could save chemical and pharmaceutical manufacturers time and money. Another advantage of the method is that, unlike previous approaches, it can be performed at temperatures and gas pressures that are easily attainable in a laboratory.

The process could also be used to alter the molecules in existing pharmaceuticals to make them more effective, safer or less addictive.

The chemists have tested their technique using very small samples of reactants—far less than a gram. But Nelson is hopeful that the methodology can be scaled up to be useful for a broad range of real-world [chemical](#) reactions.

More information: Brian Shao et al. Arylation of hydrocarbons enabled by organosilicon reagents and weakly coordinating anions, *Science* (2017). [DOI: 10.1126/science.aam7975](https://doi.org/10.1126/science.aam7975)

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