

# Efficiency of water electrolysis doubled

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Together with his team, Wolfgang Schuhmann develops new electrodes, for instance for producing hydrogen. Credit: RUB, Tim Kramer

Researchers have boosted the efficiency of water electrolysis. They applied a layer of copper atoms in a conventional platinum electrode. Thus, reaction intermediates could desorb a bit more easily from the catalyst surface. The modified system generated twice the amount of

hydrogen than a platinum electrode without a copper layer. The team from the Ruhr-Universität Bochum, Technical University in Munich and Universiteit Leiden published these results in the journal *Nature Communications*.

Water electrolysis has not yet established itself as a method for the production of hydrogen. Too much energy is lost in the process. Researchers have now doubled the efficiency of the reaction.

In the journal *Nature Communications*, researchers from the Ruhr-Universität Bochum, Technical University of Munich and Universiteit Leiden report in what way the efficiency of electrodes can be increased for the purpose of water electrolysis. Typically, platinum is applied as catalyst, in order to accelerate the conversion of water to hydrogen and oxygen. For the reaction to be as efficient as possible, intermediates must not adhere too strongly or too weakly at the catalyst surface.

## **Traditional electrodes bind intermediates too strongly**

The team headed by Prof Dr Aliaksandr Bandarenka from the Department of Physics of Energy Conversion and Storage in Munich and Prof Dr Wolfgang Schuhmann from the Center for Electrochemical Sciences in Bochum has calculated how strongly intermediates must adhere to the electrodes, in order to most efficiently facilitate the reaction. The analysis revealed that traditional electrodes from platinum, rhodium and palladium bind the intermediates a bit too strongly.

The researchers modified the properties of the platinum catalyst surface by applying a layer of [copper atoms](#). With this additional layer, the system generated twice the amount of hydrogen than with a pure platinum electrode. But only if the researchers applied the copper layer directly under the top layer of the platinum atoms. The group observed another useful side effect: the copper layer extended the service life of

the electrodes, for example by rendering them more corrosion-resistant.

## **Water electrolysis has not yet been implemented on a large scale**

Only four per cent of all hydrogen produced worldwide are the result of water electrolysis. As the [electrodes](#) used in the process are not efficient enough, large-scale application is not profitable. "To date, hydrogen has been mainly obtained from fossil fuels, with large CO<sub>2</sub> volumes being released in the process," says Wolfgang Schuhmann. "If we succeeded in obtaining [hydrogen](#) by using electrolysis instead, it would be a huge step towards climate-friendly energy conversion. For this purpose, we could utilise surplus electricity, for example generated by wind power."

"In addition, the research on this reaction allows us to test, how well we can design catalyst surfaces by precisely positioning different metal atoms," adds Aliaksandr Bandarenka. "A knowledge many other catalytic processes might benefit from."

**More information:** Jakub Tymoczko et al. Making the hydrogen evolution reaction in polymer electrolyte membrane electrolyzers even faster, *Nature Communications* (2016). [DOI: 10.1038/NCOMMS10990](https://doi.org/10.1038/NCOMMS10990)

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