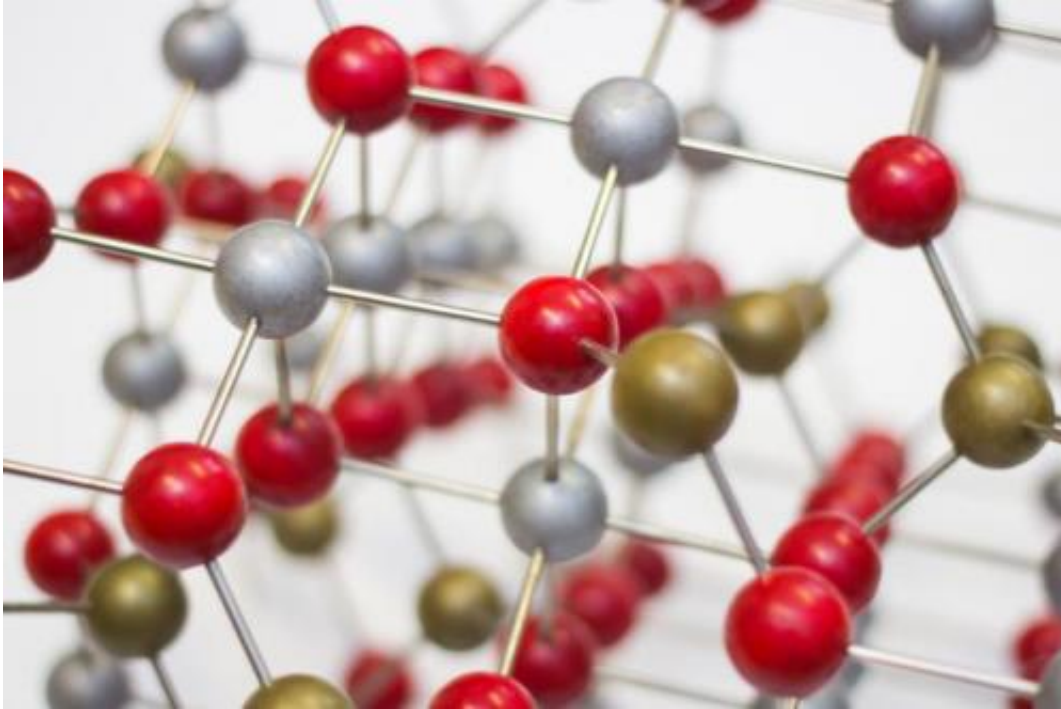


The finer details of rust

December 4 2014



The structure of iron oxide is not as simple as one might think. Credit: TU Wien

Scientists at the Vienna University of Technology have been studying the behavior of iron oxide surfaces. The atomic structure of iron oxide, which had been assumed to be well-established, turned out to be wrong. The behavior of iron oxide is governed by missing iron atoms in the atomic layer directly below the surface. This is a big surprise with potential applications in chemical catalysis, electronics or medicine.

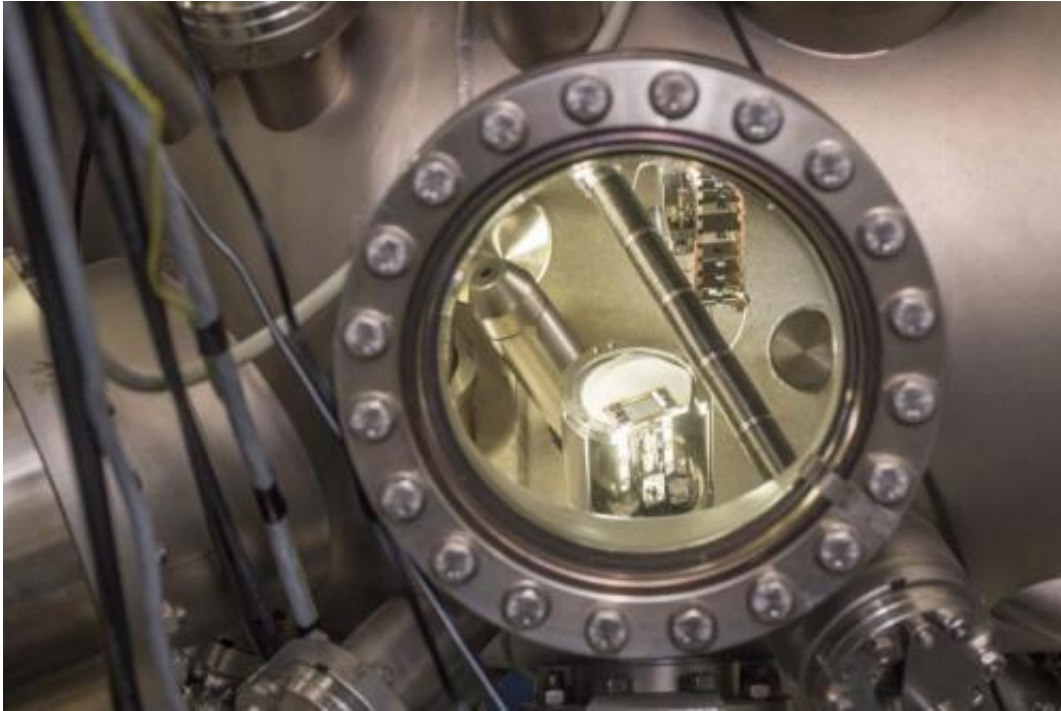
Magnetite (or Fe_3O_4) is an elaborate kind of rust - a regular lattice of

oxygen and iron atoms. But this material plays an increasingly important role as a catalyst, in electronic devices and in medical applications.

Scientists at the Vienna University of Technology have now shown that the [atomic structure](#) of the magnetite surface, which everybody had assumed to be well-established, has in fact been wrong all along. The properties of magnetite are governed by missing iron atoms in the sub-surface layer. "It turns out that the surface of Fe_3O_4 is not Fe_3O_4 at all, but rather $\text{Fe}_{11}\text{O}_{16}$ ", says Professor Ulrike Diebold, head of the metal-oxide-research group at TU Wien (Vienna). The new findings have now been published in the journal *Science*.

The Material That Just Won't Behave

Perhaps the most surprising property of the magnetite surface is that single atoms placed on the surface, for instance gold or palladium, stay perfectly in place instead of balling up and forming a nanoparticle. This effect makes the surface an extremely efficient catalyst for chemical reactions - but nobody had ever been able to tell why magnetite behaves that way. "Also, Fe_3O_4 -based electronics never function quite as well as they should", says Gareth Parkinson (TU Wien). "Because materials interact with their environment through the surface, it's really important to understand the [structure](#) of the surface and why it forms."



Iron oxide sample in the vacuum chamber. Credit: TU Wien

Very often, the properties of [metal oxides](#) depend on oxygen vacancies in the topmost atomic layers. Depending on the environment, some [oxygen atoms](#) on the surface may be missing. This can dramatically influence the electronic properties of the material. "Everyone in our community thinks about missing oxygen atoms. That is why it took us quite a while to figure out that it is in fact missing iron atoms that do the trick", says Gareth Parkinson.

It's Not the Oxygen, It's the Metal

Instead of a fixed structure of metal atoms with built-in oxygen atoms, one rather has to think of iron-oxides as a well-defined oxygen structure with little metal atoms hiding inside. Directly below the outermost atomic layer, the [crystal structure](#) is rearranged and some iron atoms are

absent.

It is precisely above such places of missing iron atoms that other [metal atoms](#) attach. These iron-vacancy-sites are regularly spaced, and so there is always some well-defined distance between gold or palladium atoms attaching to the surface. This explains why magnetite surfaces prevent these atoms from forming clusters.

The idea of completely re-thinking the crystal structure of magnetite was rather bold, and therefore the scientists analysed their theory very carefully. Quantum simulations were carried out on large supercomputers to show that the proposed structure was indeed physically reasonable. After that, electron diffraction measurements were done together with researchers at the University of Erlangen-Nuremberg, Germany.

"By scattering slow electrons at surfaces, one can measure how well the actual crystal structure of the material agrees with a theoretical model", says Ulrike Diebold. This agreement is quantified by the so-called "R-value". "For very well-known structures, one may achieve an R-value of 0.1 or 0.15. For magnetite, nobody had ever managed to get anything better than 0.3, and people said it just could not be done." But the new [magnetite](#) structure with missing [iron atoms](#) agrees very well with the experimental data, yielding an R-value of 0.125.

Magnetite Is Just the Beginning

Metal oxides are widely known to be technologically important but extremely complicated to describe. "Our results show that there is no need to despair. Metal oxides can be modelled quite accurately after all, but maybe not in the way one might expect at first glance", says Gareth Parkinson. The scientists expect that their findings do not just apply to iron oxide but also to oxides of cobalt, manganese or nickel. Re-thinking

their crystal structures could possibly boost [iron-oxide](#) research in many areas and lead to applications in chemical catalysis, electronics or medicine.

The research project is building bridges between physics and chemistry. TU Wien has set up the doctoral college "Solids4fun" to foster close interdisciplinary collaborations in the field of materials and [surface](#) science.

More information: "Subsurface cation vacancy stabilization of the magnetite (001) surface" *Science*, [www.sciencemag.org/lookup/doi/ ... 1126/science.1260556](http://www.sciencemag.org/lookup/doi/10.1126/science.1260556)

Provided by Vienna University of Technology

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