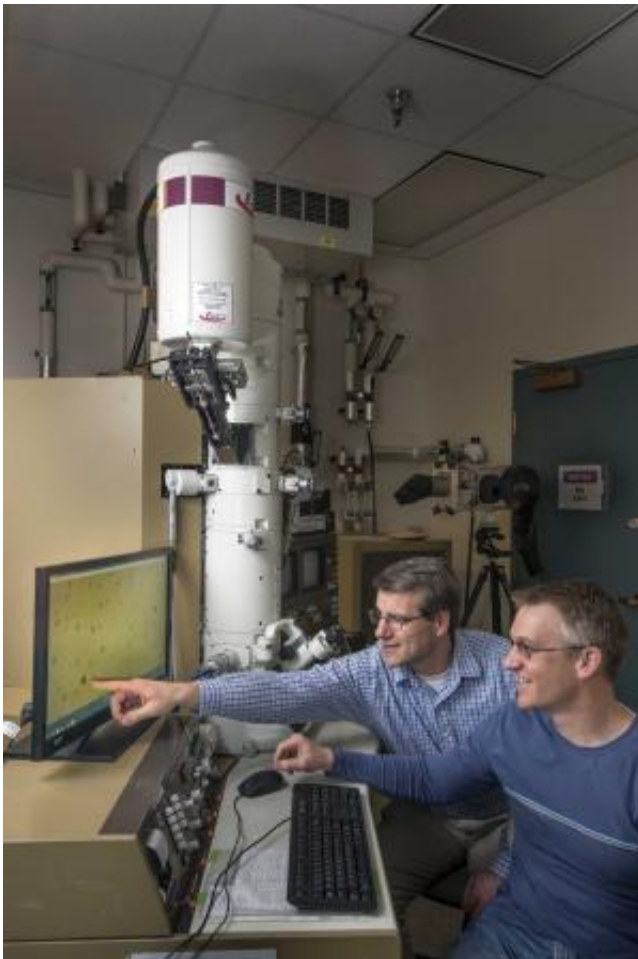


Novel nanoparticle production method could lead to better lights, lenses, solar cells

June 17 2014, by Sue Holmes



Sandia National Laboratories researchers Dale Huber, left, and Todd Monson have come up with an inexpensive way to synthesize titanium-dioxide nanoparticles, which could be used in everything from solar cells to light-emitting diodes. Credit: Randy Montoya

(Phys.org) —Sandia National Laboratories has come up with an inexpensive way to synthesize titanium-dioxide nanoparticles and is seeking partners who can demonstrate the process at industrial scale for everything from solar cells to light-emitting diodes (LEDs).

Titanium-dioxide (TiO_2) [nanoparticles](#) show great promise as fillers to tune the [refractive index](#) of anti-reflective coatings on signs and optical encapsulants for LEDs, solar cells and other optical devices. Optical encapsulants are coverings or coatings, usually made of silicone, that protect a device.

Industry has largely shunned TiO_2 nanoparticles because they've been difficult and expensive to make, and current methods produce particles that are too large.

Sandia became interested in TiO_2 for optical encapsulants because of its work on LED materials for solid-state lighting.

Current production methods for TiO_2 often require high-temperature processing or costly surfactants—molecules that bind to something to make it soluble in another material, like dish soap does with fat.

Those methods produce less-than-ideal nanoparticles that are very expensive, can vary widely in size and show significant particle clumping, called agglomeration.

Sandia's technique, on the other hand, uses readily available, low-cost materials and results in nanoparticles that are small, roughly uniform in size and don't clump.

"We wanted something that was low cost and scalable, and that made particles that were very small," said researcher Todd Monson, who along with principal investigator Dale Huber patented the process in mid-2011

as "High-yield synthesis of brookite TiO₂ nanoparticles."

Low-cost technique produces uniform nanoparticles that don't clump

Their method produces nanoparticles roughly 5 nanometers in diameter, approximately 100 times smaller than the wavelength of visible light, so there's little light scattering, Monson said.

"That's the advantage of nanoparticles—not just nanoparticles, but small nanoparticles," he said.

Scattering decreases the amount of light transmission. Less scattering also can help extract more light, in the case of an LED, or capture more light, in the case of a solar cell.

TiO₂ can increase the refractive index of materials, such as silicone in lenses or optical encapsulants. Refractive index is the ability of material to bend light. Eyeglass lenses, for example, have a high refractive index.

Practical nanoparticles must be able to handle different surfactants so they're soluble in a wide range of solvents. Different applications require different solvents for processing.

Technique can be used with different solvents

"If someone wants to use TiO₂ nanoparticles in a range of different polymers and applications, it's convenient to have your particles be suspension-stable in a wide range of solvents as well," Monson said.

"Some biological applications may require stability in aqueous-based solvents, so it could be very useful to have surfactants available that can make the particles stable in water."

The researchers came up with their synthesis technique by pooling their backgrounds—Huber's expertise in nanoparticle synthesis and polymer chemistry and Monson's knowledge of materials physics. The work was done under a Laboratory Directed Research and Development project Huber began in 2005.

"The original project goals were to investigate the basic science of nanoparticle dispersions, but when this synthesis was developed near the end of the project, the commercial applications were obvious," Huber said. The researchers subsequently refined the process to make particles easier to manufacture.

Existing synthesis methods for TiO_2 particles were too costly and difficult to scale up production. In addition, chemical suppliers ship [titanium-dioxide nanoparticles](#) dried and without surfactants, so particles clump together and are impossible to break up. "Then you no longer have the properties you want," Monson said.

The researchers tried various types of alcohol as an inexpensive solvent to see if they could get a common titanium source, titanium isopropoxide, to react with water and alcohol.

The biggest challenge, Monson said, was figuring out how to control the reaction, since adding water to titanium isopropoxide most often results in a fast reaction that produces large chunks of TiO_2 , rather than nanoparticles. "So the trick was to control the reaction by controlling the addition of water to that reaction," he said.

Textbooks said making nanoparticles couldn't be done, Sandia persisted

Some textbooks dismissed the titanium isopropoxide-water-alcohol method as a way of making TiO_2 nanoparticles. Huber and Monson, however, persisted until they discovered how to add water very slowly by

putting it into a dilute solution of alcohol. "As we tweaked the synthesis conditions, we were able to synthesize nanoparticles," Monson said.

The next step is to demonstrate synthesis at an industrial scale, which will require a commercial partner. Monson, who presented the work at Sandia's fall Science and Technology Showcase, said Sandia has received inquiries from companies interested in commercializing the technology.

"Here at Sandia we're not set up to produce the particles on a commercial scale," he said. "We want them to pick it up and run with it and start producing these on a wide enough scale to sell to the end user."

Sandia would synthesize a small number of particles, then work with a partner company to form composites and evaluate them to see if they can be used as better encapsulants for LEDs, flexible high-index refraction composites for lenses or solar concentrators. "I think it can meet quite a few needs," Monson said.

Provided by Sandia National Laboratories

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